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4.14.

$$\left(\frac{1000 \text{ g}}{1000 \text{ g}}\right) \left(\frac{1000 \text{ g}}{1000 \text{ g}}\right) = 1.000000$$

$$\text{new } \rho(\text{kg m}^{-3}) = 1.000000 \text{ kg m}^{-3} \times \left(\frac{1000 \text{ kg}}{1000 \text{ kg}}\right) = 1.000000$$

and that new number of moles of substance:

- 4.15. (a) The nuclei involved here have the same mass number 12.
 (b) The nuclei involved have the same mass number 12.
 (c) The nuclei involved have the same mass number 12.

4.16. From the data we find the molecular weight of the gas is 12.01 g mol^{-1} , the molecular weight is 12.01 g mol^{-1} .

$$\text{or } 12.01 \text{ g mol}^{-1} = 12.01 \text{ g mol}^{-1} \left(\frac{1000 \text{ g}}{1000 \text{ g}}\right) \left(\frac{1000 \text{ mol}}{1000 \text{ mol}}\right) = 1.000000$$

4.17. From the data we find $12.01 \text{ g mol}^{-1} = 12.01 \text{ g mol}^{-1}$.

$$\text{From } 12.01 \text{ g mol}^{-1} = 12.01 \text{ g mol}^{-1} \left(\frac{1000 \text{ g}}{1000 \text{ g}}\right) \left(\frac{1000 \text{ mol}}{1000 \text{ mol}}\right) = 1.000000$$

$$\text{From } 12.01 \text{ g mol}^{-1} = 12.01 \text{ g mol}^{-1} \left(\frac{1000 \text{ g}}{1000 \text{ g}}\right) \left(\frac{1000 \text{ mol}}{1000 \text{ mol}}\right) = 1.000000$$

4.18. From the data we find the molecular weight of the gas is 12.01 g mol^{-1} .

$$\text{From } 12.01 \text{ g mol}^{-1} = 12.01 \text{ g mol}^{-1} \left(\frac{1000 \text{ g}}{1000 \text{ g}}\right) \left(\frac{1000 \text{ mol}}{1000 \text{ mol}}\right) = 1.000000$$

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Jörg C. Woehl



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