

Chapter 6: Chemical Bonds Review Answer Key

- What is the most stable group on the Periodic Table? Why is this group stable?
Noble Gases (Group 18) because they have 8 valence electrons which means their outer energy level is full.
- Name the most reactive group(s) on the Periodic Table? Why are they reactive?
Alkali Metals (Group 1) & Halogens (Group 17): Alkali Metals want to lose one valence electron in order to become stable and Halogens want to gain one valence electron to be stable.

- In the boxes below draw the Electron dot diagrams for:

Sodium	Calcium	Boron	Chlorine	Krypton
Na ·	Ca · ·	· B · ·	· · : Cl · · ·	· · : Kr : · ·

- Name the three (3) types of chemical bonds:
Ionic bonds Covalent bonds Polyatomic
- Define ionic bond:
The force that holds cations and anions together.
- Define covalent bond:
A chemical bond in which two atoms share a pair of valence electrons.
- Define polyatomic ion:
A covalently bonded group of atoms that has a positive or negative charge and acts as a unit.
- In an ionic bond, electrons are transferred.
- Who are the participants in an ionic bond? **Metals and nonmetals**
- An ion with a positive charge is a cation; an ion with a negative charge is an anion.
- In a covalent bond, electrons are shared.
- Who are the participants in a covalent bond? **Two nonmetals**
- Identify whether the following compounds are covalent (two nonmetals), ionic (a nonmetal and a metal), or polyatomic (both).
 - CaCl₂ IONIC (transferred)**
 - H₂O COVALENT (shared)**
 - BaSO₄ POLYATOMIC (both)**
 - CH₄ COVALENT (shared)**
 - LiBr IONIC (transferred)**
 - NH₄Cl POLYATOMIC (both)**
 - KI IONIC (transferred)**

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Table of Contents Physical Science Chapter 6 Chemical Bonds Test Review

1. Understanding the eBook Physical Science Chapter 6 Chemical Bonds Test Review
 - The Rise of Digital Reading Physical Science Chapter 6 Chemical Bonds Test Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 6 Chemical Bonds Test Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 6 Chemical Bonds Test Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 6 Chemical Bonds Test Review
 - Personalized Recommendations
 - Physical Science Chapter 6 Chemical Bonds Test Review User Reviews and Ratings
 - Physical Science Chapter 6 Chemical Bonds Test Review and Bestseller Lists

5. Accessing Physical Science Chapter 6 Chemical Bonds Test Review Free and Paid eBooks
 - Physical Science Chapter 6 Chemical Bonds Test Review Public Domain eBooks
 - Physical Science Chapter 6 Chemical Bonds Test Review eBook Subscription Services
 - Physical Science Chapter 6 Chemical Bonds Test Review Budget-Friendly Options
6. Navigating Physical Science Chapter 6 Chemical Bonds Test Review eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 6 Chemical Bonds Test Review Compatibility with Devices
 - Physical Science Chapter 6 Chemical Bonds Test Review Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 6 Chemical Bonds Test Review
 - Highlighting and Note-Taking Physical Science Chapter 6 Chemical Bonds Test Review
 - Interactive Elements Physical Science Chapter 6 Chemical Bonds Test Review
8. Staying Engaged with Physical Science Chapter 6 Chemical Bonds Test Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 6 Chemical Bonds Test Review
9. Balancing eBooks and Physical Books Physical Science Chapter 6 Chemical Bonds Test Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 6 Chemical Bonds Test Review
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter 6 Chemical Bonds Test Review
 - Setting Reading Goals Physical Science Chapter 6 Chemical Bonds Test Review
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter 6 Chemical Bonds Test Review
 - Fact-Checking eBook Content of Physical Science Chapter 6 Chemical Bonds Test Review
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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with the alan stivell group that toured france in 1971

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