

Chapter 9

9.1 Compounds and Chemical Change

- **Atom**- the smallest unit of an element that can exist alone or in combination with other elements.
- **Compounds**- formed when atoms are held together by an attractive force called a **chemical bond**
- **Molecule**- The smallest particle of a compound or a gaseous element, that can exist and still remain the characteristic chemical properties of a substance
- **Chemical reaction**- a change in matter where **different chemical substances are created by breaking and/or forming chemical bonds.**
 - Examples: Photosynthesis
- Energy is *absorbed* -> produce new chemical substances -> more internal potential energy
- Energy is *released* when new chemical substances are produced -> less internal potential energy
- **Chemical Equation**: a way of describing what happens during a chemical reaction
 - Example: Photosynthesis equation-
Energy/Sunlight + carbon dioxide molecules + water molecules -> plant material molecules + oxygen molecules

9.2 Valence Electrons and Ions

- the number of electrons in the **outermost orbital** that **usually determines the chemical properties of an atom**
- **Valence electron**:- the number of electrons in the outermost orbital that usually determines the properties of an atom
- **Octet Rule**- after the element helium, atoms attempt to acquire an outer orbital with eight electrons through chemical reactions
-

9.3 Chemical Bonds

- **Chemical bond**-an attractive force that holds atoms together in a compound.
- Three Classes of Chemical Bonds-
 - Ionic Bonds
 - Covalent Bonds
 - Metallic Bonds
- **Ionic Bonds**- type of linkage formed from the electrostatic attraction between oppositely charged ions in a chemical compound
- **Covalent bonds**- result when atoms achieve the noble gas electron structure by *sharing* electrons. They generally formed between the nonmetallic elements on the right side of the periodic table
- **Metallic bonds**- formed in solid metals such as iron, copper, and the other metallic elements that make up about 80 percent of all the elements.
- **Heat Formation**- energy that is released during an ionic bond formation
- **Formula**- describes what elements are in the compound and in what proportions
- **Single bond**- covalent bond
- **Double bond**- formed by two pairs of electrons that are shared by two atoms

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Science California Edition Michael J. Padilla, 2007 *Leading, Teaching, and Learning the Common Core Standards* Rosemarye T. Taylor, Rebecca Watson, Joyce Nutta, 2014-07-09 Leading Teaching and Learning is a resource for teachers taking action on Common Core State Standards to enhance student learning Each chapter addresses various disciplines and grade levels with concrete examples Chapters focus on research based instruction academic language development thinking and complexity English learners non proficient readers rigor and collaboration for ongoing professional capacity building

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