According to Table 2.3. I gram of H<sub>2</sub>O gives up 2260 J when it condenses. Since each gram of ice requires 333.5 J to melt, we get

2260 J 1g 333.5 J 6.78 grams of ice that can be melted.

- 2.76. , H (prods) , H(rcts) (2 mol)(26.5kJ/mol) 0 0 53.0kJ
- 2.78. H 60, H°[CO<sub>2</sub>(g)] ,H°[C<sub>so</sub>(s)] 60, H°[O<sub>2</sub>(g)] 26367kJ/mole 60(393.5lkJ/mol) x 0 26367kJ/mole .H °[C<sub>so</sub>(s)] 2757kJ/mole
- 2.80. The reactions are:

$$\begin{split} 2[\text{NaHCO}_3(s) \ \text{Na}(s) + \frac{1}{2} \text{H}_2(g) + \text{C}(s) + \frac{3}{2} \text{O}_2(g)] & 2 \text{--}_1 H = +1901.62 \text{ kJ} \\ 2 \text{ Na}(s) + \text{C}(s) + \frac{3}{2} \text{O}_2(g) \ \text{Na}_2 \text{CO}_3(s) & _1 H = -1130.77 \text{ kJ C} \\ (s) + \text{O}_2(g) \ \text{CO}_2(g) & _1 H = -393.51 \text{ kJ} \\ \text{H}_2(g) + \frac{1}{2} \text{O}_2(g) \ \text{H}_2 \text{O}(l) & _1 H = -285.83 \text{ kJ} \end{split}$$

This yields the overall reaction (you can verify that), and the overall  $_{mn}H$  is the sum of the values on the right:  $_{mn}H = 91.51 \text{ kJ}$ .

- 2.82. The reaction is  $2 \text{ A1}(s) + \text{Fe}_2\text{O}_3(s) \text{ Al}_2\text{O}_3(s) + 2 \text{ Fe}(s)$ . The  $_{\text{ran}}H$  is (using data from the appendix): (2 mol)(0 kJ/mol) + (1 mol)(-1675.7 kJ/mol) (1 mol)(-825.5 kJ/mol) (2 mol)(0 kJ/mol) = -850.2 kJ.
- 2.84. In this case, since the combustion occurs in open atmosphere in (assumed) constant pressure, this time q<sub>0</sub> = H = -31,723 J. If the expansion of gases is done against a constant atmospheric pressure of 1 atm, to determine work we need to know the net volume of gas produced. From exercise 2.55, we found that there was a net change of -0.00491 mol of gas produced (ignoring the volume of the benzoic acid itself). Thus,
- V nRT (0.0049 mol)(0.08205 L atm/mol K)(24.6 273.15) 0.120 L less gas volume

Using this change in volume to calculate work:  $w p = V (1 \text{atm})(-0.120 \text{ L}) \frac{101.32 \text{ J}}{1 \text{ L}} \frac{12.1 \text{ J}}{\text{atm}}$ 

Finally, to calculate U:

# **Physical Chemistry David Ball Solutions Manual**

Hans Kuhn, David H. Waldeck, Horst-Dieter Försterling

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# **Table of Contents Physical Chemistry David Ball Solutions Manual**

- 1. Understanding the eBook Physical Chemistry David Ball Solutions Manual
  - The Rise of Digital Reading Physical Chemistry David Ball Solutions Manual
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Chemistry David Ball Solutions Manual
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Chemistry David Ball Solutions Manual
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Chemistry David Ball Solutions Manual
  - Personalized Recommendations
  - Physical Chemistry David Ball Solutions Manual User Reviews and Ratings
  - Physical Chemistry David Ball Solutions Manual and Bestseller Lists

- 5. Accessing Physical Chemistry David Ball Solutions Manual Free and Paid eBooks
  - Physical Chemistry David Ball Solutions Manual Public Domain eBooks
  - Physical Chemistry David Ball Solutions Manual eBook Subscription Services
  - Physical Chemistry David Ball Solutions Manual Budget-Friendly Options
- 6. Navigating Physical Chemistry David Ball Solutions Manual eBook Formats
  - o ePub, PDF, MOBI, and More
  - Physical Chemistry David Ball Solutions Manual Compatibility with Devices
  - Physical Chemistry David Ball Solutions Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Chemistry David Ball Solutions Manual
  - Highlighting and Note-Taking Physical Chemistry David Ball Solutions Manual
  - Interactive Elements Physical Chemistry David Ball Solutions Manual
- 8. Staying Engaged with Physical Chemistry David Ball Solutions Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Chemistry David Ball Solutions Manual
- 9. Balancing eBooks and Physical Books Physical Chemistry David Ball Solutions Manual
  - ∘ Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Chemistry David Ball Solutions Manual
- 10. Overcoming Reading Challenges
  - $\circ\,$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry David Ball Solutions Manual
  - Setting Reading Goals Physical Chemistry David Ball Solutions Manual
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry David Ball Solutions Manual
  - Fact-Checking eBook Content of Physical Chemistry David Ball Solutions Manual
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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