

## Balancing Equations Worksheet

- 1) \_\_\_\_\_  $\text{Na}_3\text{PO}_4$  + \_\_\_\_\_  $\text{KOH}$   $\rightarrow$  \_\_\_\_\_  $\text{NaOH}$  + \_\_\_\_\_  $\text{K}_3\text{PO}_4$
- 2) \_\_\_\_\_  $\text{MgF}_2$  + \_\_\_\_\_  $\text{Li}_2\text{CO}_3$   $\rightarrow$  \_\_\_\_\_  $\text{MgCO}_3$  + \_\_\_\_\_  $\text{LiF}$
- 3) \_\_\_\_\_  $\text{P}_4$  + \_\_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_\_  $\text{P}_2\text{O}_5$
- 4) \_\_\_\_\_  $\text{RbNO}_3$  + \_\_\_\_\_  $\text{BeF}_2$   $\rightarrow$  \_\_\_\_\_  $\text{Be}(\text{NO}_3)_2$  + \_\_\_\_\_  $\text{RbF}$
- 5) \_\_\_\_\_  $\text{AgNO}_3$  + \_\_\_\_\_  $\text{Cu}$   $\rightarrow$  \_\_\_\_\_  $\text{Cu}(\text{NO}_3)_2$  + \_\_\_\_\_  $\text{Ag}$
- 6) \_\_\_\_\_  $\text{CF}_4$  + \_\_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_\_  $\text{CBr}_4$  + \_\_\_\_\_  $\text{F}_2$
- 7) \_\_\_\_\_  $\text{HCN}$  + \_\_\_\_\_  $\text{CuSO}_4$   $\rightarrow$  \_\_\_\_\_  $\text{H}_2\text{SO}_4$  + \_\_\_\_\_  $\text{Cu}(\text{CN})_2$
- 8) \_\_\_\_\_  $\text{GaF}_3$  + \_\_\_\_\_  $\text{Cs}$   $\rightarrow$  \_\_\_\_\_  $\text{CsF}$  + \_\_\_\_\_  $\text{Ga}$
- 9) \_\_\_\_\_  $\text{BaS}$  + \_\_\_\_\_  $\text{PtF}_2$   $\rightarrow$  \_\_\_\_\_  $\text{BaF}_2$  + \_\_\_\_\_  $\text{PtS}$
- 10) \_\_\_\_\_  $\text{N}_2$  + \_\_\_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_\_\_  $\text{NH}_3$
- 11) \_\_\_\_\_  $\text{NaF}$  + \_\_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_\_  $\text{NaBr}$  + \_\_\_\_\_  $\text{F}_2$
- 12) \_\_\_\_\_  $\text{Pb}(\text{OH})_2$  + \_\_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_\_  $\text{H}_2\text{O}$  + \_\_\_\_\_  $\text{PbCl}_2$
- 13) \_\_\_\_\_  $\text{AlBr}_3$  + \_\_\_\_\_  $\text{K}_2\text{SO}_4$   $\rightarrow$  \_\_\_\_\_  $\text{KBr}$  + \_\_\_\_\_  $\text{Al}_2(\text{SO}_4)_3$
- 14) \_\_\_\_\_  $\text{CH}_4$  + \_\_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_\_  $\text{CO}_2$  + \_\_\_\_\_  $\text{H}_2\text{O}$
- 15) \_\_\_\_\_  $\text{Na}_3\text{PO}_4$  + \_\_\_\_\_  $\text{CaCl}_2$   $\rightarrow$  \_\_\_\_\_  $\text{NaCl}$  + \_\_\_\_\_  $\text{Ca}_3(\text{PO}_4)_2$
- 16) \_\_\_\_\_  $\text{K}$  + \_\_\_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_\_\_  $\text{KCl}$
- 17) \_\_\_\_\_  $\text{Al}$  + \_\_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_\_  $\text{H}_2$  + \_\_\_\_\_  $\text{AlCl}_3$
- 18) \_\_\_\_\_  $\text{N}_2$  + \_\_\_\_\_  $\text{F}_2$   $\rightarrow$  \_\_\_\_\_  $\text{NF}_3$
- 19) \_\_\_\_\_  $\text{SO}_2$  + \_\_\_\_\_  $\text{Li}_2\text{Se}$   $\rightarrow$  \_\_\_\_\_  $\text{SSe}_2$  + \_\_\_\_\_  $\text{Li}_2\text{O}$
- 20) \_\_\_\_\_  $\text{NH}_3$  + \_\_\_\_\_  $\text{H}_2\text{SO}_4$   $\rightarrow$  \_\_\_\_\_  $(\text{NH}_4)_2\text{SO}_4$

## Balancing Equations Worksheet – Answers

**Note to students:** It is acceptable to leave spaces blank when balancing equations – blank spaces are interpreted as containing the number "1".

- 1)     **1**  $\text{Na}_3\text{PO}_4$  + **3**  $\text{KOH}$   $\rightarrow$  **3**  $\text{NaOH}$  + **1**  $\text{K}_3\text{PO}_4$
- 2)     **1**  $\text{MgF}_2$  + **1**  $\text{Li}_2\text{CO}_3$   $\rightarrow$  **1**  $\text{MgCO}_3$  + **2**  $\text{LiF}$
- 3)     **1**  $\text{P}_4$  + **3**  $\text{O}_2$   $\rightarrow$  **2**  $\text{P}_2\text{O}_5$

# Physical Science Balancing Equations Answer Key

**Olga I. Agapova**



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