

Learning Outcomes

After going through this module, you are expected to:

1. Discuss key developments in the concept of the atomic structure; and
2. Give brief explanations of each of the models of atomic structure.

THE STRUCTURE OF ATOM

ACTIVITY 1 Write your answer below.

1. What is atom? _____

Most atomists like Democritus and advocates of atomic theory believed that the atom was a fundamental, indestructible, indivisible particle. This is sometimes known as the *billiard ball concept of the atom*.

ACTIVITY 2

1. Find out why it is called billiard ball concept of the atom. Write your answer on your paper briefly. (3-5 sentences only)

Robert Brown observed, under a microscope, that pollen suspended in water ejected particles that caused a jittery, irregular motion. This would later be called Brownian motion. Later on in the 20th century, this observation would be used by Albert Einstein and a French physicist, Jean Perrin, to mathematically and experimentally confirm John Dalton's Chemical Atomic Theory. This made the atom an undeniable part of how we thought about the universe. The atom is no longer debatable, the way it was in the time of the Greeks, the alchemists and Lavoisier.

FOUR ATOMIC MODELS

1. Plum-Pudding Model

Later in 1897, Joseph John Thomson published the idea that electricity was in particles that were part of the atom. Experimenting with cathode rays, he established the mass and charge properties of these particles. These particles were named electrons. In 1904, he came up with the plum-pudding model, which was an idea of what the atom looked like based on his experiments.

He would later conclude that the electron was not the only source of mass in the atom. This implied that the atom was composed of other particles.

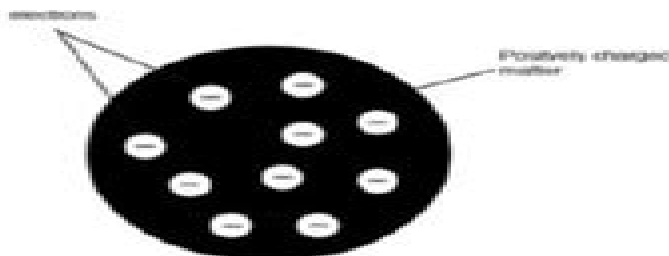


Figure 1. Thomson's plum-pudding model, a sphere with a uniformly distributed positive charge and enough embedded electrons to neutralize the positive charge. A plum pudding is a sort of cake with raisins embedded in it.

2. Rutherford's atomic Model

Ernest Rutherford, a student of Thomson's, who was among many who studied radioactivity. He concluded that radioactivity occurred due to changes on a subatomic level, or changes within the atom itself. In 1902, he worked in Thomson's laboratory where he distinguished two kinds of radiation based on their penetrating

power: α (alpha) and β (beta). He studied these types of radiation and noticed, from his experiments, that alpha particles would sometimes bounce off at a high angle when made to penetrate a very thin gold foil.

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 - The Rise of Digital Reading Physical Science Paper 1 Grade11 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Paper 1 Grade11 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Paper 1 Grade11 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Paper 1 Grade11 2014
 - Personalized Recommendations
 - Physical Science Paper 1 Grade11 2014 User Reviews and Ratings
 - Physical Science Paper 1 Grade11 2014 and Bestseller Lists
5. Accessing Physical Science Paper 1 Grade11 2014 Free and Paid eBooks

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 - Physical Science Paper 1 Grade11 2014 eBook Subscription Services
 - Physical Science Paper 1 Grade11 2014 Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Physical Science Paper 1 Grade11 2014 Compatibility with Devices
 - Physical Science Paper 1 Grade11 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Paper 1 Grade11 2014
 - Highlighting and Note-Taking Physical Science Paper 1 Grade11 2014
 - Interactive Elements Physical Science Paper 1 Grade11 2014
8. Staying Engaged with Physical Science Paper 1 Grade11 2014
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Paper 1 Grade11 2014
9. Balancing eBooks and Physical Books Physical Science Paper 1 Grade11 2014
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Paper 1 Grade11 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Paper 1 Grade11 2014
- Setting Reading Goals Physical Science Paper 1 Grade11 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Paper 1 Grade11 2014
- Fact-Checking eBook Content of Physical Science Paper 1 Grade11 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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