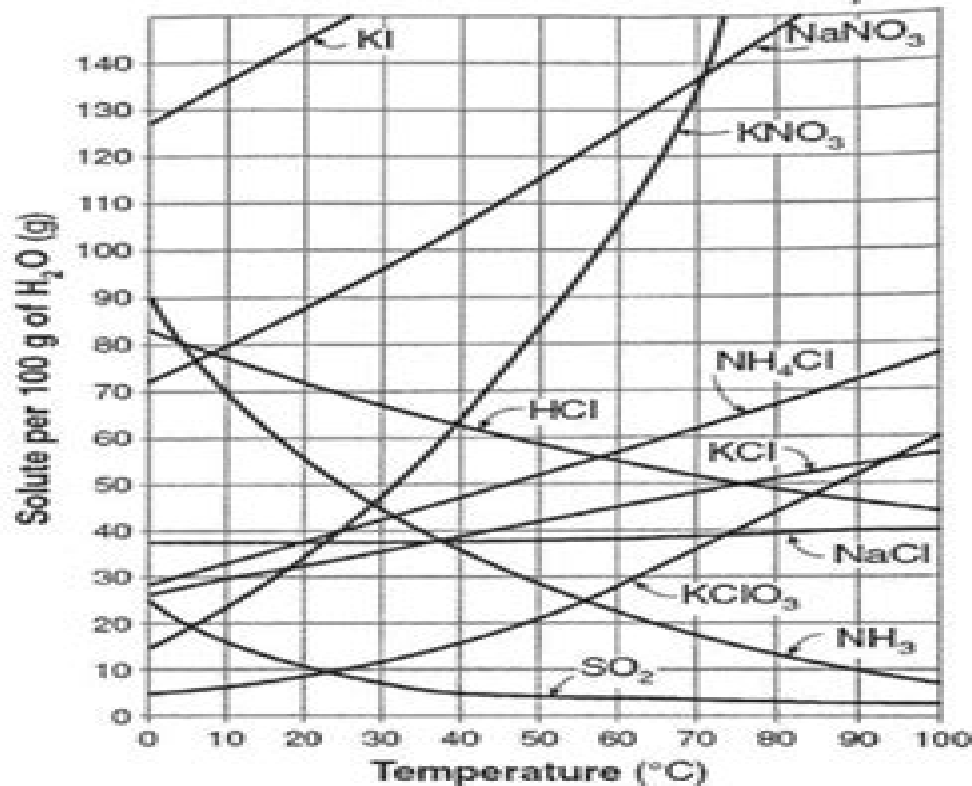


Reading a Solubility Chart

- 1) The curve shows the # of grams of solute in a saturated solution containing 100 mL or 100 g of water at a certain temperature.
- 2) Any amount of solute below the line indicates the solution is unsaturated at a certain temperature
- 3) Any amount of solute above the line in which all of the solute has dissolved shows the solution is supersaturated.
- 4) If the amount of solute is above the line but has not all dissolved, the solution is saturated and the # grams of solute settled on the bottom of the container = total # g in solution - # g of a saturated solution at that temperature. (according to the curve)
- 5) Solutes whose curves move upward w/ increased temperature are typically solids b/c the solubility of solids increases w/ increased temperature.
- 6) Solutes whose curves move downward w/ increased temperature are typically gases b/c the solubility of gases decreases with increased temperature.



Solubility Problems

1. Which of the salts shown on the graph is the most soluble in water at 10°C? **KI**
2. At what temperature is 90g of NaNO₃ a saturated solution in 100grams of water? **≈ 23°C**
3. A solution has 70grams of KNO₃ and is unsaturated, what is the minimum temperature? **46°C**
4. At 20°C, a saturated solution of KClO₃ is created in water. How many grams of potassium chlorate must be added to saturate the solution at 50°C? **12g**
5. Which of the salts shown on the graph is the least soluble in water at 10°C? **KClO₃**
6. Which of the salts shown on the graph has the greatest increase in solubility as the temperature increases from 30 degrees to 60 degrees? **KNO₃**

Reading A Solubility Chart Answers

Tracy Lynn Haroff

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Reading A Solubility Chart Answers:

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Partial Molal Properties this edition uses the more popular Gibbs Energy and Partial Molar Properties changes in symbols the first edition used the Lewis Randal fugacity rule and the popular symbol for the same quantity this edition only uses the popular notation and new problems have been added to the text Finally the second edition includes an appendix about the Bridgman table and its use

Cracking the ACT, 2010 Edition Princeton Review, Geoff Martz, 2009-11 A guide to preparing for the ACT based on the Princeton Review coaching course offers advice on test taking specific suggestions for each section of the exam and three full length practice exams

Improving Student Comprehension in Chemistry Laboratories Tracy Lynn Haroff, 2006

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The Journal of Industrial and Engineering Chemistry, 1914

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