

CHAPTER 8 STUDY GUIDE

Covalent Bonding

Section 8.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond

molecule

sigma bond

exothermic

pi bond

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1)** _____. When such an attachment is formed, bond dissociation energy is released, and the process is **(2)** _____. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3)** _____. If the electrons shared are centered between the two atoms, the attachment is called a(n) **(4)** _____. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5)** _____.

In your textbook, read about single and multiple bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

6. In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
 - a. as single atoms
 - b. as molecules containing two atoms
 - c. as molecules containing three atoms
 - d. as molecules containing four atoms
7. How many electrons are shared in a double covalent bond?
 - a. none
 - b. one
 - c. two
 - d. four
8. Bond length is the distance between
 - a. two molecules of the same substance.
 - b. the electrons in two attached atoms.
 - c. the nuclei of two attached atoms.
 - d. the orbitals of two attached atoms.
9. Which of the following relationships relating to bond length is generally correct?
 - a. the shorter the bond, the stronger the bond
 - b. the shorter the bond, the weaker the bond
 - c. the shorter the bond, the fewer the electrons in it
 - d. the shorter the bond, the lower the bond dissociation energy

Reading For 8 2 The Nature Of Covalent Bonding

IM Harris



Reading For 8 2 The Nature Of Covalent Bonding:

Symmetry and Physical Properties of Crystals Cécile Malgrange, Christian Ricolleau, Michel Schlenker, 2014-12-04

Crystals are everywhere from natural crystals minerals through the semiconductors and magnetic materials in electronic devices and computers or piezoelectric resonators at the heart of our quartz watches to electro optical devices Understanding them in depth is essential both for pure research and for their applications This book provides a clear thorough presentation of their symmetry both at the microscopic space group level and the macroscopic point group level The implications of the symmetry of crystals for their physical properties are then presented together with their mathematical description in terms of tensors The conditions on the symmetry of a crystal for a given property to exist then become clear as does the symmetry of the property The geometrical representation of tensor quantities or properties is presented and its use in determining important relationships emphasized An original feature of this book is that most chapters include exercises with complete solutions This allows readers to test and improve their understanding of the material The intended readership includes undergraduate and graduate students in materials science and materials related aspects of electrical and optical engineering researchers involved in the investigation of the physical properties of crystals and the design of applications based on crystal properties such as piezoelectricity electro optics optical activity and all those involved in the characterization of the structural properties of materials **University of Michigan Official Publication**

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The Nature of the Chemical Bond and the Structure of Molecules and Crystals Linus Pauling,1960 Thorough discussion of the various types of bonds their relative natures and the structure of molecules and crystals

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