

C11- PHYSICAL SCIENCE
 Third Monthly Test Reviewer

Coverage: **MODULE 1**
 UNIT 1: Atomic Theory
 UNIT 2: The Discovery and Synthesis of Chemical Elements

UNIT 1: Atomic Theory

- Atoms are composed of **protons, neutrons, and electrons.**



Table 2.3 Properties of Subatomic Particles

Particle	Symbol	Mass (kg)	Relative Mass (proton =1)	Relative Charge
Proton	p^+	1.673×10^{-27}	1	+1
Neutron	n^0	1.675×10^{-27}	1	0
Electron	e^-	9.109×10^{-31}	0.00055	-1

- An atom is mostly empty space with a highly dense, **positive** center called **nucleus**.
- For a neutral atom, the **number of protons is equivalent to the number of electrons.**
(Lithium has 3 protons therefore it also has 3 electrons)
- The **atomic number** of an element is also equivalent to its number of **protons.**
(Nitrogen has an atomic number of 7 therefore it has 7 protons in its nucleus)
- The **atomic mass** (mass number) is primarily determined by the number of protons and **neutrons.**

- Isotopes** are atoms with the **same atomic number** but different mass numbers. They **differ in the number of neutrons.**
(Carbon-12, Carbon-13, Carbon-14)
- Ion** is an element or atom that has a **positive or negative net charge.** It is charged because the number of electrons does not equal the number of protons in the atom.

Two types:

- Cation**- positive ion (loses electron)
- Anion**- negative ion (gains electron)



- Take note that when an **atom is not electrically neutral**, it means that it does not have the same number of protons and electrons.
- When an atom or element is **positively charged**, it means that it has **more protons** in its nucleus.

Historical Development of Atom

- Democritus & Leucippus:**
Atomism: All matter is made up of very small, indivisible particles called atoms (atomos).
- Aristotle:** Four elements in nature (air, fire, water, and earth)
- John Dalton: Billiard Ball Model/Solid Sphere Model:**
Atomic Theory: All matter is made up of very small, indivisible particles called atoms.

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