

Physical Chemistry II – Exam 1 SOLUTIONS

Constant	Value in SI Units
h	$6.62607 \times 10^{-34} \text{ J s}$
$\hbar = h/2\pi$	$1.05457 \times 10^{-34} \text{ J s}$
c	$2.99793 \times 10^8 \text{ m/s}$
e	$1.60217 \times 10^{-19} \text{ C}$
m_e	$9.10938 \times 10^{-31} \text{ kg}$
ϵ_0	$8.85419 \times 10^{-12} \text{ C}^2 \text{J}^{-1} \text{m}^{-1}$
R (Rydberg const.)	109737.3 cm^{-1}

- 1.) (17 points) When light with a wavelength of 3000 \AA shines on a metal surface, the kinetic energy of ejected electrons is $2.960 \times 10^{-19} \text{ J}$. When light with a wavelength of 5400 \AA shines on the same metal surface, the kinetic energy of ejected electrons is $1.362 \times 10^{-21} \text{ J}$. From this information, determine the value of Planck's constant (in units of J s). [$1 \text{ \AA} = 10^{-10} \text{ m}$]

For the photoelectric effect, $KE = h\nu - \phi$. From this equation, we can see that a plot with kinetic energy on the y -axis and frequency on the x -axis yields a straight line with the slope equal to Planck's constant h . Therefore, one way to determine Planck's constant from the data given is to calculate the slope,

$$h = \text{slope} = \frac{\Delta KE}{\Delta \nu}$$

The kinetic energy values are given in the problem, but the frequencies must be calculated from the wavelengths.

$$\begin{aligned} \nu_1 &= \frac{c}{\lambda_1} \\ &= \frac{2.99793 \times 10^8 \text{ m/s}}{(3000 \text{ \AA}) \left(\frac{10^{-10} \text{ m}}{1 \text{ \AA}} \right)} \\ \nu_1 &= 9.993 \times 10^{14} \text{ s}^{-1} \end{aligned}$$

$$\begin{aligned} \nu_2 &= \frac{c}{\lambda_2} \\ &= \frac{2.99793 \times 10^8 \text{ m/s}}{(5400 \text{ \AA}) \left(\frac{10^{-10} \text{ m}}{1 \text{ \AA}} \right)} \\ \nu_2 &= 5.552 \times 10^{14} \text{ s}^{-1} \end{aligned}$$

Physical Chemical Self Test Solution

J.J. Moore



Physical Chemical Self Test Solution:

Student Solutions Manual for Physical Chemistry C. A. Trapp, Peter Atkins, Julio de Paula, 2009-12-18 With its modern emphasis on the molecular view of physical chemistry its wealth of contemporary applications vivid full color presentation and dynamic new media tools the thoroughly revised new edition is again the most modern most effective full length textbook available for the physical chemistry classroom Available in Split Volumes For maximum flexibility in your physical chemistry course this text is now offered as a traditional text or in two volumes Volume 1 Thermodynamics and Kinetics ISBN 1 4292 3127 0 Volume 2 Quantum Chemistry Spectroscopy and Statistical Thermodynamics ISBN 1 4292 3126 2 Solutions Manual to Accompany Physical Chemistry for the Life Sciences C. A. Trapp, M. P. Cady, 2011 This solutions manual contains fully worked solutions to all end of chapter discussion questions and exercises featured in Physical Chemistry for the Life Sciences *Students Solutions Manual to Accompany Physical Chemistry: Quanta, Matter, and Change 2e* Charles Trapp, Marshall Cady, Carmen Giunta, 2014 The Students Solutions Manual to Accompany Physical Chemistry Quanta Matter and Change 2e provides full worked solutions to the exercises and the odd numbered discussion questions and problems presented in the parent book The manual is intended for students and instructors alike and provides helpful comments and friendly advice to aid understanding Elements of Physical Chemistry Peter Atkins, Julio de Paula, 2013 Elements of Physical Chemistry has been carefully crafted to help students increase their confidence when using physics and mathematics to answer fundamental questions about the structure of molecules how chemical reactions take place and why materials behave the way they do Atkins' Physical Chemistry Peter William Atkins, Julio de Paula, 2014-03 PART 1 THERMODYNAMICS PART 2 STRUCTURE PART 3 CHANGE **Physical Chemistry** Peter Atkins, Julio de Paula, Ronald Friedman, 2014 This title takes an innovative molecular approach to the teaching of physical chemistry The authors present the subject in a rigorous but accessible manner allowing students to gain a thorough understanding of physical chemistry **Chemical Structure and Reactivity** James Keeler, Peter Wothers, 2013-11 Why do certain substances react together in the way that they do What determines the shape of molecules And how can we predict whether a particular reaction will happen at all Such questions lie at the heart of chemistry the science of understanding the composition of substances their reactions and properties Though introductory chemistry is often broken into three sections inorganic organic and physical the only way for students to fully understand the subject is to see it as a single unified whole Chemical Structure and Reactivity rises to the challenge of depicting the reality of chemistry Offering a fresh approach to the subject by depicting it as a seamless discipline the text shows how organic inorganic and physical concepts can be blended together in order to achieve the common goal of understanding chemical systems With a lively and engaging writing style enhanced by vivid illustrations only Chemical Structure and Reactivity makes teaching chemistry with an integrated approach possible Special Features The only introductory text to take a truly integrated approach in explaining the fundamentals of chemistry Fosters

an orbital based understanding of reactions with clear curly arrow mechanistic detail throughout A two part structure allows flexibility of use Part I lays down the core of the subject while Part II describes a series of relatively standalone topics which can be selected to fit a particular course Numerous concepts are illustrated with fully cross referenced custom developed online modules enabling students to develop an understanding through active learning Self test exercises embedded in the text with solutions at the end of each chapter and extensive question sets encourage hands on learning to help students master the subject and gain confidence The Online Resource Centre features a range of additional resources for both students and registered adopters of the book New to this Edition A new chapter on symmetry has been added to Part I Discussions of organometallic chemistry spectroscopy and molecular geometry have been expanded Cross references from Part I to Part II have been increased to make the links between core concepts and more advanced topics clearer More self test questions and exercises have been provided

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Physical Chemistry for the Life Sciences Peter Atkins,Julio de Paula,2011 Peter Atkins and Julio de Paula offer a fully integrated approach to the study of physical chemistry and biology

US Solutions Manual to Accompany Elements of Physical Chemistry 7e David Smith,2017-09-28 The Solutions Manual to Accompany Elements of Physical Chemistry 7th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions

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Elements of Physical Chemistry Peter William Atkins,Julio De Paula,2017 This revision of the introductory textbook of physical chemistry has been designed to broaden its appeal particularly to students with an interest in biological applications

Solutions Manual to Accompany Inorganic Chemistry Alen Hadzovic,2018 As you master each chapter in Inorganic Chemistry having detailed solutions handy allows you to confirm your answers and develop your ability to think through the problem solving process

Atkins' Physical Chemistry 11e Peter Atkins,Julio De Paula,James Keeler,2019-09-06 Atkins Physical Chemistry Molecular Thermodynamics and Kinetics is designed for use on the second semester of a quantum first physical chemistry course Based on the hugely popular Atkins Physical Chemistry this volume approaches molecular

thermodynamics with the assumption that students will have studied quantum mechanics in their first semester The exceptional quality of previous editions has been built upon to make this new edition of Atkins Physical Chemistry even more closely suited to the needs of both lecturers and students Re organised into discrete topics the text is more flexible to teach from and more readable for students Now in its eleventh edition the text has been enhanced with additional learning features and maths support to demonstrate the absolute centrality of mathematics to physical chemistry Increasing the digestibility of the text in this new approach the reader is brought to a question then the math is used to show how it can be answered and progress made The expanded and redistributed maths support also includes new Chemist s toolkits which provide students with succinct reminders of mathematical concepts and techniques right where they need them Checklists of key concepts at the end of each topic add to the extensive learning support provided throughout the book to reinforce the main take home messages in each section The coupling of the broad coverage of the subject with a structure and use of pedagogy that is even more innovative will ensure Atkins Physical Chemistry remains the textbook of choice for studying physical chemistry

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Chemical Metallurgy J.J. Moore,2013-10-22 Chemical Metallurgy Second Edition provides the fundamental chemical principles and demonstrates the application of these principles to process metallurgy materials synthesis and processing and corrosion protection The book consists of nine chapters The first five chapters emphasize the fundamental chemical principles involved in metallurgical reactions An additional chapter on slag chemistry has also been added in this second edition in order to provide a more thorough understanding of slag metal reactions The final three chapters focus on the applications of the chemical principles to the extraction and refining of metals metal melting and recycling and metallic corrosion The book will be of value to materials students and teachers and scientists and engineers entering employment in the metallurgical and materials processing and metal finishing industries [Research Grants Index](#) National Institutes of Health (U.S.). Division of Research Grants,1967

Chemical Principles Peter Atkins,Loretta Jones,2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry s frontier continually demonstrating how to solve problems think about nature and

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