

Chapter 1

Gases and the Zeroth Law of Thermodynamics

- 1.2.** A system is any part of the universe under observation. The “surroundings” includes everything else in the universe. Consider a solution calorimeter in which two aqueous solutions are mixed and temperature changes are recorded. In this case the solutes (reactants and products) would be considered the “system”. The water, calorimeter, and rest of the lab and world would be the “surroundings”.
- 1.4.** (a) $12.56 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 12,560 \text{ cm}^3 = 1.256 \times 10^4 \text{ cm}^3$
(b) $45^\circ\text{C} + 273.15 = 318 \text{ K}$
(c) $1.055 \text{ atm} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} \times \frac{100,000 \text{ Pa}}{1 \text{ bar}} = 1.069 \times 10^5 \text{ Pa}$
(d) $1233 \text{ mmHg} \times \frac{1 \text{ torr}}{1 \text{ mmHg}} \times \frac{1 \text{ atm}}{760 \text{ torr}} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} = 1.644 \text{ bar}$
(e) $125 \text{ mL} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 125 \text{ cm}^3$
(f) $4.2 \text{ K} - 273.15 = -269.0^\circ\text{C}$
(g) $25750 \text{ Pa} \times \frac{1 \text{ bar}}{100,000 \text{ Pa}} = 0.2575 \text{ bar}$
- 1.6.** $p_{\text{tot}} = p_{\text{atmosph}} + \rho hg$, where ρhg correspond to the pressure exerted by the liquid
 $p_{\text{tot}} - p_{\text{atmosph}} = \rho hg = (1.0 \times 10^3 \text{ kg/m}^3)(0.23 \text{ m})(9.80 \text{ m/s}^2) = 2254 \text{ N/m}^2 = 2.3 \times 10^3 \text{ Pa}$
- 1.8.** In terms of the zeroth law of thermodynamics, heat will flow from the (hot) burner or flame on the stove into the (cold) water, which gets hotter. Then heat will move from the hot water into the (colder) egg.
- 1.10.** For this sample of gas under these conditions, $F(T) = 2.97 \text{ L} \times 0.0553 \text{ atm} = 0.164 \text{ L} \cdot \text{atm}$. If the pressure were increased to 1.00 atm: $0.164 \text{ L} \cdot \text{atm} = (1.00 \text{ atm}) \times V$; therefore $V = 0.164 \text{ L}$.
- 1.12.** $V = R \cdot \frac{nT}{p}$, which rearranges to $R = \frac{pV}{nT}$.
Therefore: $R = \frac{(2.66 \text{ bar})(27.5 \text{ L})}{(1.887 \text{ mol})(466.9 \text{ K})} = 0.0830 \frac{\text{L} \cdot \text{bar}}{\text{mol} \cdot \text{K}}$

Physical Chemistry 2 Edition Solution Manual

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Table of Contents Physical Chemistry 2 Edition Solution Manual

1. Understanding the eBook Physical Chemistry 2 Edition Solution Manual
 - The Rise of Digital Reading Physical Chemistry 2 Edition Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry 2 Edition Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry 2 Edition Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry 2 Edition Solution Manual
 - Personalized Recommendations
 - Physical Chemistry 2 Edition Solution Manual User Reviews and Ratings
 - Physical Chemistry 2 Edition Solution Manual and Bestseller Lists
5. Accessing Physical Chemistry 2 Edition Solution Manual Free and Paid eBooks
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 - Highlighting and Note-Taking Physical Chemistry 2 Edition Solution Manual
 - Interactive Elements Physical Chemistry 2 Edition Solution Manual
- 8. Staying Engaged with Physical Chemistry 2 Edition Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry 2 Edition Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry 2 Edition Solution Manual
 - Setting Reading Goals Physical Chemistry 2 Edition Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry 2 Edition Solution Manual
 - Fact-Checking eBook Content of Physical Chemistry 2 Edition Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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