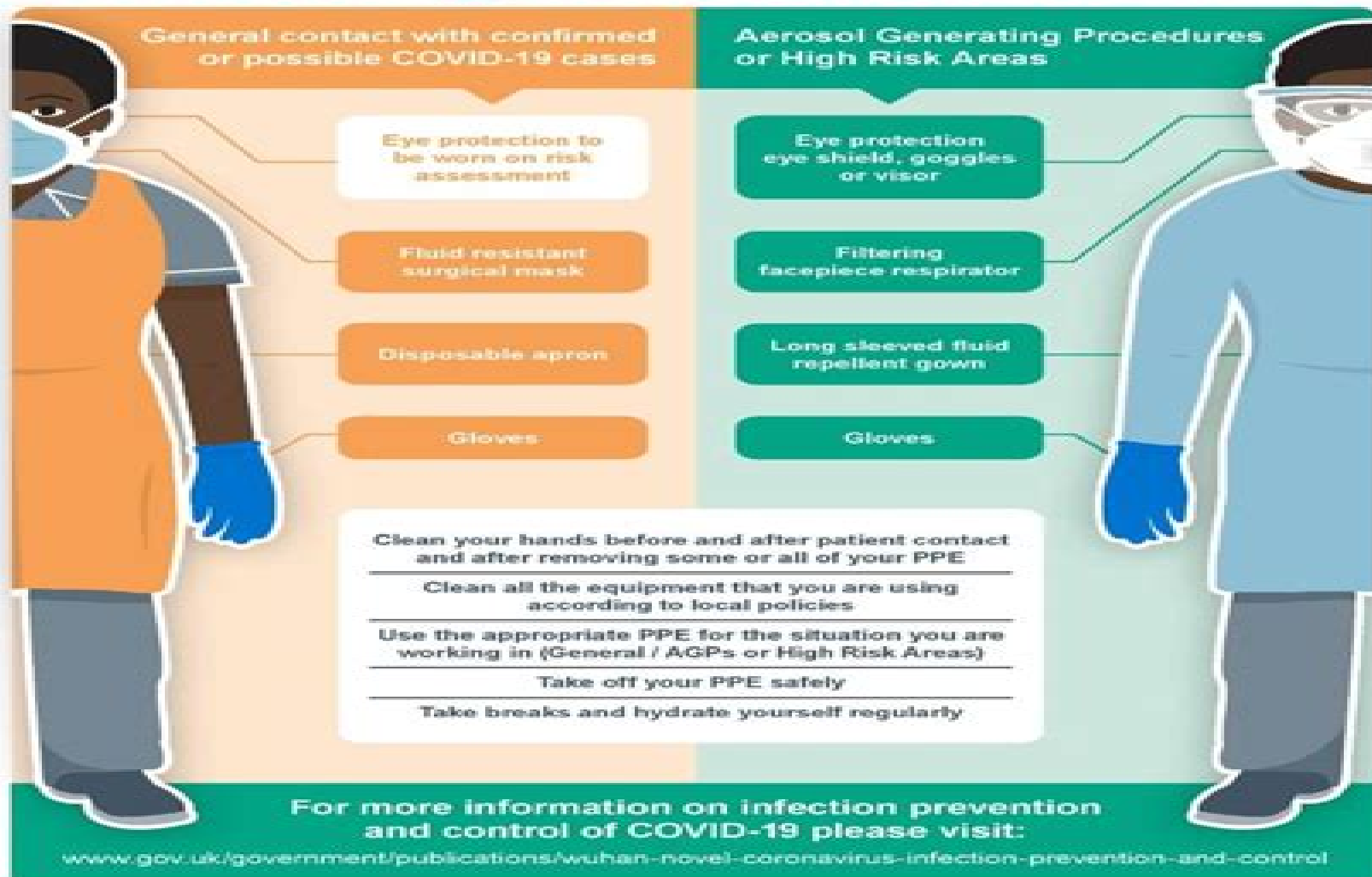


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web equilibrium moles $HSCN$ $2.00 \times 10^{-5} \text{ mol}$ $3.00 \times 10^{-6} \text{ mol}$ $1.70 \times 10^{-5} \text{ mol}$ $HSCN$ knowing the number of moles of Fe^{3+} and

hscn present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species
 $\text{mol Fe}^{3+} = 1.70 \times 10^{-5} \text{ mol}$
 $\text{mol Fe}^{3+} = 8.50 \times 10^{-4} \text{ mol}$

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reactants remain constant if we run a reaction in a closed system so that the products cannot escape we often find the reaction does not give a 100 yield of products

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