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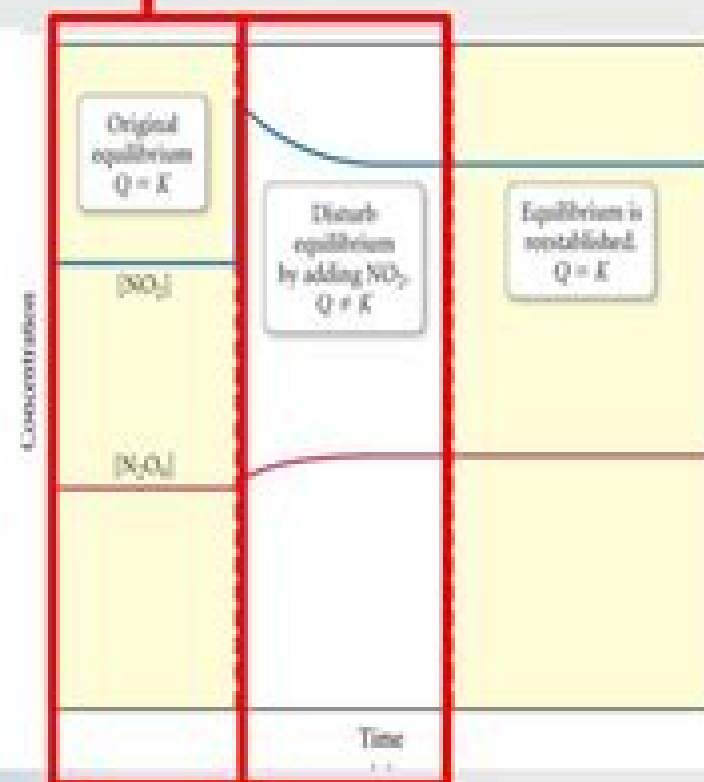
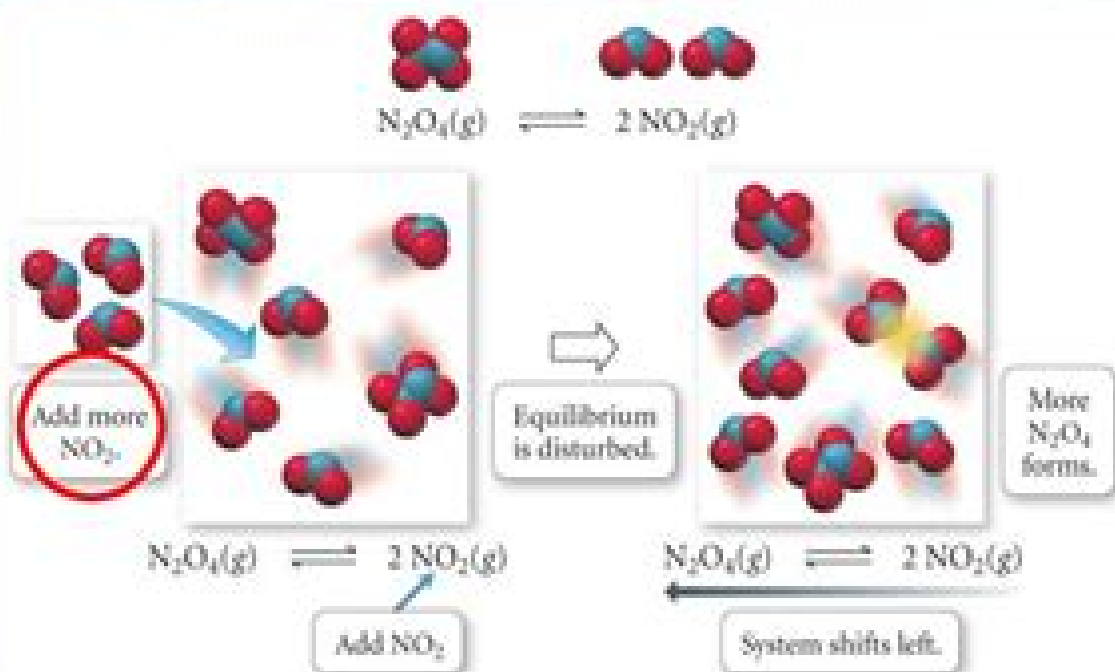
$$Q = \frac{[N_2O_4]_{eq}}{([NO_2]_{eq} + [NO_2])^2} < K$$

rate forward = rate backward

$$k_f[N_2O_4]_{eq} = k_b[NO_2]_{eq}^2$$

rate forward < rate backward

$$k_f[N_2O_4]_{eq} < k_b([NO_2]_{eq} + [NO_2])^2$$



Practical Report Effect Of Concentration On Equilibrium

Karin Nielsen-Saines



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