

### Statistics Review

1. The combustion of propene occurs via the reaction:  
$$\text{C}_3\text{H}_6\text{O(g)} + 5\text{O}_2\text{(g)} \rightarrow 3\text{CO}_2\text{(g)} + 3\text{H}_2\text{O(g)}$$
  
How many grams of oxygen are required to burn completely 50.0 g of propene?
2. Small amounts of ethane can be prepared in the laboratory by the reaction:  
$$\text{CH}_3\text{I(g)} + \text{Al(CH}_3)_3\text{(g)} \rightarrow \text{CH}_4\text{(g)} + \text{Al(CH}_3)_2\text{(g)}$$
  
What volume of ethane at 0°C and 1 atm can be prepared from 10.0 g of methyl iodide?
3. The heat content per unit volume is important. Find the heat content (kJ/mol) of ethane. How many grams of ethane are required to produce 1000 g of water according to the following reaction?  
$$2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(g)}$$
4. Ethanol is oxidized to ethanoic acid by chromic acid. The overall reaction that takes place is:  
$$\text{C}_2\text{H}_5\text{OH(g)} + 2\text{O}_2\text{(g)} \rightarrow \text{C}_2\text{H}_3\text{O}_2\text{(g)} + 3\text{H}_2\text{O(g)}$$
  
Calculate the volume of oxygen at STP required to convert 100 g of ethanol to ethanoic acid and water.
5. Acetylene, which is widely used as a fuel, can be prepared from acetylene hydrate (calcium acetylide) and oxygen. It is used for welding (acetylene torches) and also as a feedstock for organic synthesis. The equation for the reaction is:  
$$2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$$
  
How many kilograms of acetylene hydrate will be produced from 1000 kg of calcium acetylide and 500 kg of water available for the reaction?
6. Ethanol can be prepared by adding ethylene to concentrated solutions of sulfuric acid. How many grams of sulfuric acid are formed if 100 g of ethylene reacts with an excess of sulfuric acid according to the reaction below. Ethanol formed is 100% pure and there is no loss of water. How many grams of sulfuric acid are required?
7. How many kilograms of 100% sulfuric acid will be required to react with 1000 g of ethyl alcohol according to the following equation:  
$$\text{C}_2\text{H}_5\text{OH} + \text{H}_2\text{SO}_4 \rightarrow \text{C}_2\text{H}_5\text{OSO}_3\text{H} + \text{H}_2\text{O}$$
8. 10.0 g of sulfuric acid is reacted with an excess of ethane. Calculate the mass of phosphorus pentachloride produced.
9. How many grams of hydrogen are produced by the decomposition of 100 g of calcium hydroxide?
10. 100 g of oxygen gas at STP can be prepared by the decomposition of 1000 g of potassium chlorate?
11. How many grams of zinc are required for the reduction of 1000 g of hydrogen bromide acid?
12. What mass of phosphorus chloride is required to react with 100 g of sulfur dioxide?

# Review Chemistry 1 Honors Unit 10 Stoichiometry

**Ying-Ying Zheng**



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