Shill furview (Shill - Student will accurately describe the relationship between P. V. T and n of garees)

In your everyday life, you often expensed the different gas laws discussed in class, in action. In the table below are some situations that may be familiar to you. Think about each situation and in column 2, describe in your own words what is happening on the microscopic (molecular) level, in column 3, write the name of the simple law (Avogadro's, Boyle's, Charles' or Dalton's) that best describes the situation, in the two extra rows at the bottom, add some additional situations, not stready discussed, that you have expenienced.

Everyday Situation	Microscopic Description	Laver
The pressure shown on a tire gauge doubles as the mass of arris doubled (@ a constant temperature)		Avogadro's
2. A ballicon over the mouth of a bottle containing air begins to expand when it is moved in to sunlight.		Charles
An car piston compresses gases while the engine is working.		Dayle's
An inflated rubber raft gets softer when some of the gas is allowed to escape.	1	Avogadro's
5. A balloon placed in a sauna increases in size.		Charles"
6. A hot ar balloon lands when burners under the balloon's open end turn off		Chades
7. When you aqueste an inflated balloon, it seems to push back harder in to your hand.		Doyse's
A highly pressured tank of helium gas will fill hundreds of balloons.		Doyle's
9. Red. yellow and blue bimpo balls are flying around inside a bingo-caller's dispenser. The effect on the sides of the container is the same as if an equal number of all red bingo balls are in the container.		Distron's
 When the number of bings balls is doubled, the effect on the sides of the container is doubled. 		Avegadro's
11		

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Arturo Cuomo

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