

**Advance Study Assignment:** Resolution of Matter into Pure Substances.  
II. Fractional Crystallization

1. Using Figure 3.1, determine

- a. the number of grams of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  that will dissolve in 100 g of  $\text{H}_2\text{O}$  at  $100^\circ\text{C}$ .

180 g  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

- b. the number of grams of water required to dissolve 4.0 g of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  at  $100^\circ\text{C}$ . (Hint: your answer to Problem 1(a) gives you the needed conversion factor for g  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  to g  $\text{H}_2\text{O}$ .)

At  $100^\circ\text{C}$ , grams of water = (given mass of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ )

At  $100^\circ\text{C}$ , grams of water (4.0  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ )  $\times$  (100g  $\text{H}_2\text{O}$  / 180g  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ )

2.22 g  $\text{H}_2\text{O}$

At  $100^\circ\text{C}$ , grams of water = 2.22 g

- c. the number of grams of water required to dissolve 18 g  $\text{KNO}_3$  at  $100^\circ\text{C}$ .

From the given figure, at  $100^\circ\text{C}$ , 240 g of  $\text{KNO}_3$  is soluble in 100 g water. To dissolve 18 g  $\text{KNO}_3$ :

(18.0g  $\text{KNO}_3$ )  $\cdot$  (100g  $\text{H}_2\text{O}$  / 240g  $\text{KNO}_3$ )

7.50 g  $\text{H}_2\text{O}$

at  $100^\circ\text{C}$  grams of water = 2.50g

- d. the number of grams of water required to dissolve a mixture containing 18 g  $\text{KNO}_3$  and 4.0 g  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ , assuming that the solubility of one substance is not affected by the presence of another.

At  $100^\circ\text{C}$ , 2.22g of water is required to dissolve 4.0 g  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  and 7.5 g of water is required to dissolve 18.0 g  $\text{KNO}_3$

Grams of water required to dissolve 4.0g of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  and 18.0 g of  $\text{KNO}_3$  at  $100^\circ\text{C}$  is 7.5 g

# Resolution Of Matter Into Pure Substances Answers

**Anthony S. Fauci**



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