SS 2 CHEMISTRY PRACTICAL 3RD TERM EXMINATION

 B is a solution containing 5.30g of anhydrous sodium trioxocarbonate (IV) per dm³. A is a solution of hydrochloric acid of unknown concentration.

Put A into the burette and titrate 25cm* (or 20cm*) portions of 8 with A. Using methyl orange as the indicator. Record the volume of your pipette.

Tabulate your burette readings, and calculate the average volume of acid used.

From your results, calculate:

- a. The concentration of solution B in moles per dm1
- b. The concentration of solution A in moles per dm3
- The mass concentration in g per dm³ of the acid in solution A.

The equation of reaction is:

$$2HCI+Na_2CO_3 \rightarrow 2NaCI+CO_2+H_2O$$

- (a). If you were provided with anhydrous CaCO₃, spatula and stirrer (i) list three other materials
 you would require to prepare a standard solution of CaCO₃ (ii) State what you would observe on
 adding dilute HCl to a portion of the CaCO₃
 - (b). State what would be observed when aqueous ammonia is added to: (i) Litmus paper; (ii) Pb(NO₁)₂ solution in drops until in excess.

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Nigel S. Key, Michael Makris, David Lillicrap

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- 1. Understanding the eBook Ssschemistry Practical 3rd Term
 - The Rise of Digital Reading Ssschemistry Practical 3rd Term
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Ssschemistry Practical 3rd Term
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ssschemistry Practical 3rd Term
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ssschemistry Practical 3rd Term
 - Personalized Recommendations
 - Ssschemistry Practical 3rd Term User Reviews and Ratings
 - Ssschemistry Practical 3rd Term and Bestseller Lists
- 5. Accessing Ssschemistry Practical 3rd Term Free and Paid eBooks

- Ssschemistry Practical 3rd Term Public Domain eBooks
- Ssschemistry Practical 3rd Term eBook Subscription Services
- Ssschemistry Practical 3rd Term Budget-Friendly Options
- 6. Navigating Ssschemistry Practical 3rd Term eBook Formats
 - o ePub, PDF, MOBI, and More
 - Ssschemistry Practical 3rd Term Compatibility with Devices
 - Ssschemistry Practical 3rd Term Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ssschemistry Practical 3rd Term
 - Highlighting and Note-Taking Ssschemistry Practical 3rd Term
 - Interactive Elements Ssschemistry Practical 3rd Term
- 8. Staying Engaged with Ssschemistry Practical 3rd Term
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ssschemistry Practical 3rd Term
- 9. Balancing eBooks and Physical Books Ssschemistry Practical 3rd Term
 - Benefits of a Digital Library
 - $\circ\,$ Creating a Diverse Reading Collection Ssschemistry Practical 3rd Term
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ssschemistry Practical 3rd Term
 - Setting Reading Goals Ssschemistry Practical 3rd Term
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ssschemistry Practical 3rd Term
 - Fact-Checking eBook Content of Ssschemistry Practical 3rd Term
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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