



Underlined words and phrases are to be filled in by students on the Note-taking Worksheet.

Section 1 Stability in Bonding

- A. Some elements combine chemically and no longer have the same properties they did before forming a compound.
- B. A(n) chemical formula is composed of symbols and subscripts indicating the number of atoms of an element in a compound.
- C. Atoms form compounds when the compound is more stable than the separate atoms.
1. Noble gases are more chemically stable than other elements because they have a complete outer energy level.
 2. Elements that do not have full outer energy levels are more stable in compounds.
 3. Atoms can lose, gain, or share electrons to get a stable outer energy level.
 4. A(n) chemical bond is the force that holds atoms together in a compound.

DISCUSSION QUESTION:

Why do elements form compounds? *To become more chemically stable by getting a complete outer energy level*

Section 2 Types of Bonds

- A. A(n) ion is a charged particle because it has more or fewer electrons than protons.
1. When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge.
 2. When an atom gains an electron, it becomes a negatively charged ion.
- B. An ionic compound is held together by the ionic bond—the force of attraction between opposite charges of the ions.
1. The result of this bond is a(n) neutral compound.
 2. The sum of the charges on the ions is zero.
- C. Molecules are neutral particles formed as a result of sharing electrons.
1. A covalent bond is the force of attraction between atoms sharing electrons.
 2. Atoms can form double or triple bonds depending on whether they share two or three pairs of electrons.

Stability In Bonding Section 1 Study Guide

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is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Study Guide & Solutions Manual to Accompany Organic Chemistry, Third Edition G. Marc Loudon, Joseph G. Stowell, 1995

Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Katherine J. Denniston, 2006

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Organic Chemistry Study Guide & Problems Book Brent L. Iverson, 1998 Companion to Organic chemistry 2nd ed by William H Brown Provides a detailed section by section overview of the major points covered in the text All of the problems from the text are reprinted here with detailed stepwise solutions

Study Guide for Chang's Chemistry Kenneth W. Watkins, 1988 *Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition* Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry John McMurry, Susan McMurry, 1986

Study Guide for Chemical Principles Thomas Elliott Taylor, 1979 *Student Study Guide to accompany Chemistry* Martin Silberberg, Libby Weberg, 2005-01-06

Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide

includes many new worked out examples

The United States Catalog Mary Burnham, Carol Hurd, 1928

Stability In Bonding Section 1 Study Guide Book Review: Unveiling the Magic of Language

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