

### THIRD TERM EXAMINATION QUESTION

#### CHEMISTRY: SSS II

1. Consider the following reaction equation:  $C_2H_{6g} + 3O_{2g} \rightarrow 2CO_{2g} + 2H_2O_{lg}$ . The volume of oxygen at s.t.p. that will be required to burn 14g of ethane is? [ $C_2H_6 = 28$ , Molar volume of gas at s.t.p. =  $22.4 \text{ dm}^3$ ] (a)  $67.2 \text{ dm}^3$  (b)  $33.6 \text{ dm}^3$  (c)  $11.2 \text{ dm}^3$  (d)  $3.73 \text{ dm}^3$
2.  $50\text{cm}^3$  of sulphur (IV) oxide,  $800\text{cm}^3$  of ammonia,  $450\text{cm}^3$  of hydrogen chloride,  $1.0\text{cm}^3$  of carbon (IV) oxide will respectively saturate  $1.0\text{cm}^3$  of water at  $15^\circ\text{C}$ . Which of the following is suitable for demonstrating the fountain experiment? (a) Sulphur (IV) oxide and hydrogen chloride (b) carbon (IV) oxide and ammonia (c) Ammonia and hydrogen chloride (d) carbon (IV) oxide and sulphur (IV) oxide
3. The volume of  $0.25 \text{ mol dm}^{-3}$  solution of KOH that would yield  $6.5\text{g}$  of solid KOH on evaporation is? [ $K = 39$ ,  $O = 16$ ,  $H = 1$ ] (a)  $464.30 \text{ cm}^3$  (b)  $625 \text{ cm}^3$  (c)  $1000 \text{ cm}^3$  (d)  $2153.80 \text{ cm}^3$
4. Solubility is practically applied in (a) fractional distillation (b) solvent extraction (c) the determination of pH (d) the determination of saturation in hydrocarbons
5. The following factors affect the solubility of a solid in a given solvent **EXCEPT** (a) nature of solute (b) nature of solvent (c) pressure (d) temperature



- Which of the following deductions could be **correctly** made from the graph? (a) the solubility of U is not affected by change in temperature (b) the solubility of S decreases with increasing temperature (c) T is most soluble among the salts (d) S is least soluble among the salts
7. Soap lather in is an example of a colloid in which a (a) liquid is dispersed in gas (b) solid is dispersed in liquid (c) gas is dispersed in liquid (d) liquid is dispersed in liquid
  8. The number of sulphur atoms in  $3.20 \text{ g}$  of  $SO_{2g}$  is? [ $O = 16$ ,  $S = 32$ , Avogadro's constant =  $6.02 \times 10^{23}$ ] (a)  $3.01 \times 10^{22}$  (b)  $6.02 \times 10^{22}$  (c)  $6.02 \times 10^{23}$  (d)  $1.20 \times 10^{24}$
  9. At  $27^\circ\text{C}$ ,  $58.5\text{g}$  of sodium chloride is present in  $250\text{cm}^3$  of a solution. The solubility of sodium chloride at this temperature is? [Molar mass of sodium chloride =  $111.0\text{g mol}^{-1}$ ] (a)  $2.11 \text{ mol dm}^{-3}$  (b)  $0.25 \text{ mol dm}^{-3}$  (c)  $1.00 \text{ mol dm}^{-3}$  (d)  $0.50 \text{ mol dm}^{-3}$
  10. Which of the following pollutants will lead to the depletion of ozone layer? (a) chlorofluorocarbon (b) carbon (II) oxide (c) hydrogen sulphide (d) sulphur (IV) oxide
  11. Consider the equation:  $Cr_2O_7^{2-} + 6Fe^{2+} + 14H^+ \rightarrow 2Cr^{3+} + 6Fe^{3+} + 7H_2O$ . The oxidation number of chromium changes from (a)  $+5$  to  $+3$  (b)  $+6$  to  $+3$  (c)  $-2$  to  $+3$  (d)  $+7$  to  $+3$

# Ss1 Test Question On Chemistry

**Jianjun Gao**

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## **Table of Contents Ss1 Test Question On Chemistry**

1. Understanding the eBook Ss1 Test Question On Chemistry
  - The Rise of Digital Reading Ss1 Test Question On Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Ss1 Test Question On Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ss1 Test Question On Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ss1 Test Question On Chemistry
  - Personalized Recommendations
  - Ss1 Test Question On Chemistry User Reviews and Ratings
  - Ss1 Test Question On Chemistry and Bestseller Lists
5. Accessing Ss1 Test Question On Chemistry Free and Paid eBooks

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- Ss1 Test Question On Chemistry Budget-Friendly Options
- 6. Navigating Ss1 Test Question On Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Ss1 Test Question On Chemistry Compatibility with Devices
  - Ss1 Test Question On Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ss1 Test Question On Chemistry
  - Highlighting and Note-Taking Ss1 Test Question On Chemistry
  - Interactive Elements Ss1 Test Question On Chemistry
- 8. Staying Engaged with Ss1 Test Question On Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ss1 Test Question On Chemistry
- 9. Balancing eBooks and Physical Books Ss1 Test Question On Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ss1 Test Question On Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ss1 Test Question On Chemistry
  - Setting Reading Goals Ss1 Test Question On Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ss1 Test Question On Chemistry
  - Fact-Checking eBook Content of Ss1 Test Question On Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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