

Valence Electron Worksheet

Lewis Structures

Name _____

Date _____ Period _____

1. How many electrons are in the following elements?

A. Carbon

C. Indium

E. Tungsten

B. Chlorine

D. Calcium

F. Antimony

2. Draw Lewis Structures for each of the following elements. How many electrons are in their outer shells?

A. Silicon

C. Sulfur

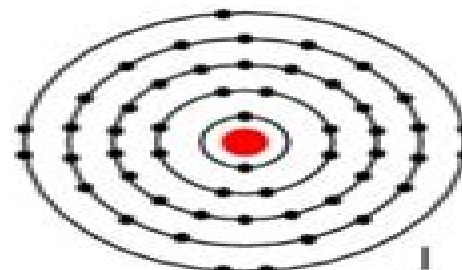
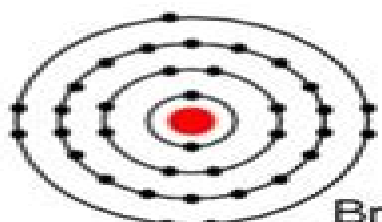
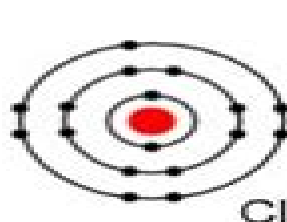
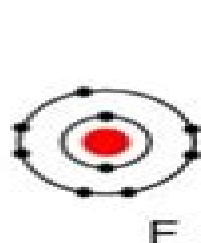
E. Beryllium

B. Potassium

D. Boron

F. Argon

3. Circle the valence electrons only. How many valence electrons does each atom have?



4. Finish the sentence: Elements in the same group have the same number of _____ electrons.

5. Fill in the electrons for the following elements. Use your work to answer how many valence electrons each element has.

Example: C (carbon): 1S $\uparrow\downarrow$ 2S $\uparrow\downarrow$ 2P \uparrow \uparrow \uparrow 3S _____ 3P _____ #valence electrons 4

He: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in helium _____

Cl: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in chlorine _____

Li: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in lithium _____

B: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in boron _____

F: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in fluorine _____

Ar: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in argon _____

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