



(a) Sodium metal



(b) Chlorine gas



(c) Sodium metal reacting with chlorine gas



(d) Sodium chloride is formed

Figure 31: Reaction of sodium metal with chlorine.

(iii) Magnesium on heating with chlorine readily forms magnesium chloride, which is an ionic chloride :



(iv) Aluminium reacts with chlorine, on heating, to form aluminium chloride :



(v) Zinc combines directly with chlorine to form zinc chloride :



(vi) Iron combines with chlorine, when heated, to form iron (III) chloride :



(vii) On heating, copper reacts with chlorine to form copper (II) chloride :



All these metal chlorides are ionic compounds (or electrovalent compounds).

6. Reaction of Metals with Hydrogen

Metals generally do not react with hydrogen because metals form compounds by losing electrons (which are accepted by other elements) and hydrogen also forms compounds by losing electron (or by sharing of electrons). So, normally a hydrogen atom does not accept the electrons given by a metal atom to form a compound. But a few very reactive metals (like sodium, potassium, calcium and magnesium) can force the hydrogen atoms to accept electrons given by them and form salt-like solid compounds called metal hydrides. Thus, most of the metals do not combine with hydrogen. Only a few reactive metals like sodium, potassium, calcium and magnesium react with hydrogen to form metal hydrides. Metal hydrides are ionic compounds formed by the transfer of electrons from metal atoms to hydrogen atoms. In a metal hydride, the hydrogen is present in the form of a negative ion (anion) called hydride ion, H^- . Here are some examples.

(i) When hydrogen gas is passed over heated sodium, then sodium hydride is formed :



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