

SSA GUIDELINES**Practical work:**

- Learners do TWO experiments (ONE Chemistry, ONE Physics) for SSA. Choose three experiments from this A.T.P.
- Term 1: Choose ONE experiment. Record in term one.
- Term 2 OR Term 3: Choose ONE experiment. Record the mark in term 2.
- The suggested formal experiments NOT chosen for SSA, should become informal experiments.

2020: Grade 12 Term 1: PHYSICAL SCIENCES

TERM 1 (32 days)	Week 1 15/1 – 17/1 (3 days)	Week 2 20/1 – 24/1 (5 days)	Week 3 27/1 – 31/1 (5 days)	Week 4 3/2 – 7/2 (5 days)	Week 5 10/2 – 14/2 (5 days)	Week 6 17/2 – 21/2 (5 days)	Week 7 24/2 – 28/2 (5 days)	Week 8 3/3 – 7/3 (5 days)	Week 9-10 10/3 – 21/3 (12 days)	Week 11 24/3 – 28/3 (5 days)
CAPS Topics	MECHANICS: Momentum & Impulse (2 hrs)	MECHANICS: Momentum & Impulse (2 hrs)	MECHANICS: Vertical projectile motion (2 hrs)	MECHANICS: Vertical projectile motion (2 hrs)	WAVES, SOUND & LIGHT: Doppler Effect (2 hrs)	MATTER 1: MATERIALS: Organic molecules (2 hrs)	MATTER 1: MATERIALS: Organic molecules (2 hrs)	MATTER 1: MATERIALS: Organic molecules (2 hrs)	MATTER 1: MATERIALS: Organic molecules (2 hrs)	CONTROL TEST (2 hrs)
Topics / Concepts, Skills and Values	<ul style="list-style-type: none"> Define & calculate the momentum of a moving object $p = mv$ Describe the vector nature of momentum & draw vector diagrams. State Newton's second law in terms of momentum: $F_{net} = \frac{\Delta p}{\Delta t}$ Calculate the change in momentum when a resultant force acts on an object. 	<ul style="list-style-type: none"> Define impulse Use the impulse-momentum theorem ($F_{net} \Delta t = \Delta p$) in calculations for a variety of situations (one dimension) Impulse and safety considerations State the principle of conservation of linear momentum Explain what is meant by an isolated system, internal and external forces. Apply conservation of momentum to collisions (one dimension). Distinguish between elastic and inelastic collisions by calculation. 	<ul style="list-style-type: none"> Explain what is meant by a projectile. Use equations of motion to determine the position, velocity and displacement of a projectile at any given time. Sketch v vs t, a vs t and s vs t graphs for a free falling object, an object thrown vertically upwards, object thrown vertically downwards & bouncing objects. Use equations for position versus time and velocity versus time for the graphs of 1D projectile motion. 	<ul style="list-style-type: none"> For given v vs t, a vs t or s vs t graphs, determine position, displacement and velocity or acceleration at any time t. For given s vs t, v vs t or a vs t graphs, describe the motion of an object (starting, thrown vertically upwards & thrown vertically downwards and so on). 	<ul style="list-style-type: none"> State the Doppler effect and explain using illustrations the change in pitch observed when a source moves toward or away from a listener (sound and ultraviolet). State applications of the Doppler effect (Example ultrasound in medicine) State problems using $f_o = \frac{v}{\lambda_o}$, $f_s = \frac{v}{\lambda_s}$, $f_o = \frac{v}{\lambda_o \pm \frac{v_o}{\cos \theta}}$, $f_s = \frac{v}{\lambda_s \pm \frac{v_s}{\cos \theta}}$ when $\theta = 0^\circ$ or 180° Sketch source or listener waves. State light, explain red shifts & use the Doppler effect to explain why we conclude that the universe is expanding. 	<ul style="list-style-type: none"> Define organic molecules, functional groups, hydrocarbon, homogeneous series, saturated, unsaturated, and structural isomer. Write condensed, structural & molecular formulae (over 8 C atoms, 1 functional group per molecule) for alkanes (no rings), alkenes (no rings), alkynes, alcohols, halo alkanes (no rings), carboxylic acids, aldehydes, ketones, esters. Write IUPAC names for structural condensed structural formulae for compounds from above series 	<ul style="list-style-type: none"> Write IUPAC names from structural or condensed structural formulae for compounds (over 10 C atoms) (one functional group per molecule, max. two functional groups for halo alkanes) Identify other substituents (halogen and alkyl) max. 10000 alkyl substituents. Identify compounds that are saturated, unsaturated, structural isomers, alkenes, alkenes, alkyne, alcohols, aldehydes, ketones, esters, carboxylic acids, amines, nitriles. Physical properties, boiling point, melting point, vapour pressure 	<ul style="list-style-type: none"> Relationship between physical properties and strength of MP, type of functional group, chain length and branching Combustion of alkanes in excess oxygen and use in fuels Equation & reaction conditions for the formation of an ester and IUPAC names for reactant and products Chemically reactions as elimination, addition or substitution. Equations and reaction conditions for addition reactions of alkenes 	<ul style="list-style-type: none"> Equations and reaction conditions for elimination reactions: dehydrohalogenation of halo alkanes, cracking of alkanes, dehydration of alcohols. Equations and reaction conditions for substitution reactions: hydrolysis of halo alkanes, saponification of esters 	CONTROL TEST (2 hrs) Discussion and retelling work on test (2hrs) Revision (2 hours) ONE PAPER (100 marks 2 hours) : Physics <ul style="list-style-type: none"> Newton's laws of motion Momentum and impulse Vertical projectile motion Doppler Effect 40 marks : Chemistry <ul style="list-style-type: none"> Organic molecules Discussion and retelling work of control test.
Requires pre-knowledge	Newton's laws of motion	Newton's laws of motion Equations of motion	Equations of motion	Equations of motion	<ul style="list-style-type: none"> Wave properties: frequency, wavelength, amplitude Graphical representation of waves 	Chemical bonding Valency	Chemical bonding Valency	Hydrocarbon series	IUPAC naming, writing chemical formulae	SSA

Sacai June 2014 Physical Science

Lei Shi



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