

Review for Thermochemistry

Vocabulary: exothermic, endothermic, activation energy, net energy, $q = mc\Delta T$, $mc\Delta T = -mc\Delta T$, potential energy of reactants, potential energy of products, heat of reaction ΔH , specific heat.

Calculations Practice:

You WILL NOT be told which equation to use on the exam. The two equations you will need to be able to calculate are $q = mc\Delta T$ and $mc\Delta T = -mc\Delta T$. REMEMBER: if you are putting a metal into water, then you use $mc\Delta T = -mc\Delta T$ and if you are only dealing with one thing, use $q = mc\Delta T$.

- Find the specific heat of an unknown metal if adding 12 grams of 100°C sample to 100.0 grams of 25°C water raises the final temperature of both objects to 28°C . $mc\Delta T = -mc\Delta T$

metal

$$m = 12$$

$$c = ?$$

$$\Delta T = -72^\circ\text{C}$$

$$T_f = 28^\circ\text{C}$$

$$T_i = 100^\circ\text{C}$$

$$mc\Delta T = -mc\Delta T$$

$$12(c)(-72) = -100(4.18)(3)$$

$$\frac{-864c}{-864} = \frac{-1254}{-864}$$

$$c = 1.45$$

$$\boxed{c = 1.45}$$

H₂O

$$m = 100\text{g}$$

$$c = 4.18\text{J/g}\cdot^\circ\text{C}$$

$$\Delta T = 3^\circ\text{C}$$

$$T_f = 28^\circ\text{C}$$

$$T_i = 25^\circ\text{C}$$

- Calculate the heat gained or lost when 25 grams of water are heated from 15°C to 100°C .

$$q = mc\Delta T$$

$$Q = ?$$

Prentice Hall Thermochemistry Section Review Answer Key

J Dewey



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of the committee on Society history issued from time to time beginning with v 30 Feb 1908 **Chemical Age** ,1957
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