

## Lab Report

## Aspirin Synthesis

### A. Preparation Data:

1. Mass of salicylic acid (g)

3.0028 g

### B. Purification of Aspirin Data:

1. Mass of filter paper (g)

0.3200 g

2. Mass of filter paper & aspirin (g)

0.4284 g

3. Mass of aspirin (g)

0.1048 g ← Actual yield

4. Temperature when melting begins

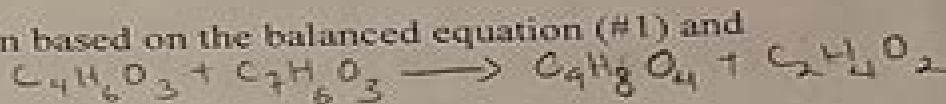
100.23°C Initial

5. Temperature when melting is complete

100.40°C Final

### Results and Considerations:

1. Calculate the theoretical yield of aspirin based on the balanced equation (#1) and your starting mass of salicylic acid.



2. Calculate the percent yield for your experiment.

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \quad \% =$$

3. If the yield of your aspirin is greater than 100%, what must you do experimentally to obtain a more reasonable yield?

# Preparation Of Aspirin Lab Report Answers

**Jo Allan Beran**



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General, Organic, and Biological Chemistry Study Guide and Selected Solutions Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems     Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section     **Experimental Organic Chemistry** Daniel R. Palleros, 2000-02-04 This cutting edge lab manual takes a multiscale approach presenting both micro semi micro and macroscale techniques The manual is easy to navigate with all relevant techniques found as they are needed Cutting edge subjects such as HPLC bioorganic chemistry multistep synthesis and more are presented in a clear and engaging fashion     Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science     **Laboratory Experiments for General Chemistry** Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998     **Laboratory Exercises for Preparatory Chemistry** Kathy Dodds Tyner, 1994-06 Laboratory Exercises for Preparatory Chemistry is the perfect complement to a one semester preparatory chemistry laboratory course Tyner's manual

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