

Preparation of Solution

Molar Solutions

What do you know?

Mass of Solute (g)

Final Volume of Solution (L)

$$\text{Mass of Solute (g)} = \frac{1}{\text{Molar Mass } \left(\frac{\text{g}}{\text{mole}}\right)} \times \frac{1}{\text{Final Volume of solution (L)}} \times \text{Molarity (M or } \frac{\text{mole}}{\text{L}})$$

What do you want to know?

Molarity or Molar Solution (M)

$$\text{Molarity } \left(\frac{\text{mole}}{\text{L}}\right) = \frac{\text{Mass of Solute (g)}}{\text{Final Volume of solution (L)} \times \text{Molar Mass } \left(\frac{\text{g}}{\text{mole}}\right)} = \text{Mass of Solute (g)}$$

What do you know?

Molarity or Solution (M)

Final Volume of Solution (L)

What do you want to know?

Mass of Solute (g)

$$\text{Mass of Solute (g)} = \frac{1}{\text{Molar Mass } \left(\frac{\text{g}}{\text{mole}}\right)} \times \frac{1}{\text{Molarity (M or } \frac{\text{mole}}{\text{L}})} \times \text{Final Volume of Solution (L)}$$

What do you know?

Mass of Solute (g)

Molarity or Solution (M)

What do you want to know?

Final Volume of Solution (L)

Percent (%) Solutions

Mass/Mass Percent Concentration Solution - All percentages should be in decimal form

What do you know?

Mass of Solute (g)

Mass of Solvent (g)

$$\frac{\text{Mass of Solute (g)}}{\text{Mass of Solute (g)} + \text{Mass of Solvent (g)}} = \text{Percent Concentration of Solution (\%)}$$

What do you want to know?

Percent Concentration of Solution (%)

What do you know?

Mass of Solute (g)

Percent Concentration of Solution (%)

$$\frac{\text{Mass of Solute (g)}}{\text{Percent Concentration of Solution (\%)}} = \text{Mass of Solvent (g)} + \text{Mass of Solute (g)}$$

What do you want to know?

Mass of Solvent (g)

What do you know?

Mass of Solvent (g)

Percent Concentration of Solution (%)

$$\frac{\text{Percent Concentration of Solution (\%)}}{1 - \text{Percent Concentration of Solution (\%)}} = \text{Mass of Solute (g)}$$

What do you want to know?

Mass of Solute (g)

Mass/Volume Percent Solutions - All percentages should be in decimal form

What do you know?

Mass of Solute (g)

Final Volume of Solution (mL)

$$\frac{\text{Mass of Solute (g)}}{\text{Final Volume of Solution (mL)}} = \text{Percent Concentration of Solution (\%)}$$

What do you want to know?

Percent Concentration of Solution (%)

What do you know?

Mass of Solute (g)

Percent Concentration of Solution (%)

$$\frac{\text{Mass of Solute (g)}}{\text{Percent Concentration of Solution (\%)}} = \text{Final Volume of Solution (mL)}$$

What do you want to know?

Final Volume of Solution (mL)

What do you know?

Final Volume of Solution (mL)

Percent Concentration of Solution (%)

$$\text{Final Volume of Solution (mL)} = \frac{\text{Mass of Solute (g)}}{\text{Percent Concentration of Solution (\%)}}$$

What do you want to know?

Mass of Solute (g)

Dilutions

Molar Solution:

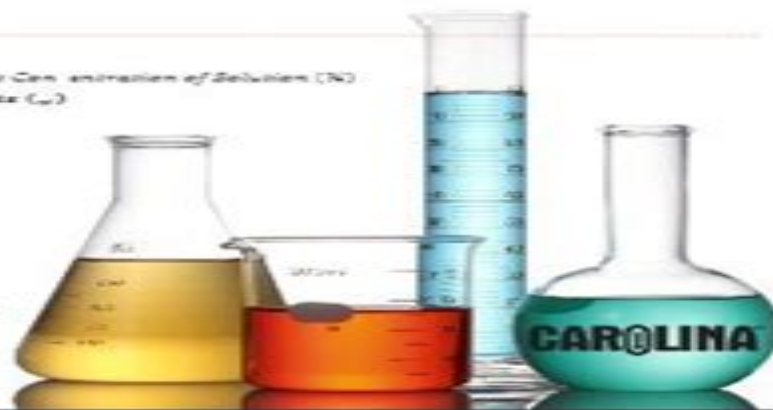
Molarity of Initial Solution $\left(\frac{\text{mole}}{\text{L}}\right) \times \text{Volume of Initial Solution (L)} =$

Molarity of Final Solution $\left(\frac{\text{mole}}{\text{L}}\right) \times \text{Volume of Final Solution (L)}$

Percent Concentration Solutions

Percent Concentration of Initial Solution $\left(\frac{\text{g}}{\text{mL}}\right) \times \text{Volume of Initial Solution (L)} =$

Percent Concentration of Final Solution $\left(\frac{\text{g}}{\text{mL}}\right) \times \text{Volume of Final Solution (L)}$



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Pamela Mills,Amina El-Ashmawy



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Proceedings of the American Pharmaceutical Association at the Annual Meeting American Pharmaceutical Association. Annual Meeting, American Pharmaceutical Association, 1898 Vols for 1853 1911 include list of members

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International Trade Robert C. Feenstra, Alan M. Taylor, 2008 Combining classic international economics with straight from the headlines immediacy Feenstra and Taylor s text seamlessly integrates the subject s established core content with new topic areas and new ideas that have emerged from recent empirical studies Like no other textbook it brings cutting edge theory evidence and policy analysis to the field of international economics International Economics is available as a complete textbook or in two split volumes International

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biochemistry will find this book highly informative **Intermediate Accounting** Donald E. Kieso, Jerry J. Weygandt, Terry D. Warfield, 2010-10-04 Due to the emergence of IFRS as the required convention for reporting to stock exchanges in the European Union and other important markets accountants must gain a strong understanding of these standards Intermediate Accounting integrates this new information throughout the chapters so they ll learn how to apply the new global accounting standards Global examples are presented to clearly show how the information is utilized in the field The use of various currencies is also explored which is critical for accountants to know in today s global businesses environment

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