

D 40059

(Pages : 3)

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SIXTH SEMESTER B.Sc. DEGREE EXAMINATION, MARCH/APRIL 2018

(CUCBCSS—UG)

Chemistry

CHE 6B 11—PHYSICAL CHEMISTRY—III

Time : Three Hours

Maximum : 80 Marks

Section A

*Answer in one word or sentence.*

*Answer all questions.*

1. Write the equation relating equivalent conductance and concentration of a strong electrolyte.
2. Define electrochemical series.
3. What is a fuel cell ?
4. Write the Nernst equation for electrode potential.
5. Why is it necessary to use a salt bridge in a Galvanic cell ?
6. Define buffer capacity.
7. Why does specific conductance of a solution decreases with dilution ?
8. What is meant by imperfection or defect in a crystal ?
9. Which colligative property is preferred for the molar mass determination of macromolecules ?
10. Calculate the Miller indices for crystal planes with intercepts 2a, 1b, 3c.

(10 × 1 = 10 marks)

Section B

*Answer any ten questions.*

*Each question carries 2 marks.*

11. How will you determine ionic product of water by conductance measurement ?
12. Explain the Lewis concepts of acids and bases.
13. Write down the electrode reaction and cell reaction in the following cell :  
$$\text{Pt}, \text{H}_2(\text{g}), \text{H}^+(\text{aq}) // \text{Cl}^-(\text{aq}), \text{AgCl}, \text{Ag}.$$
14. Write two advantages of  $\text{H}_2\text{-O}_2$  fuel cell over ordinary cell.
15. What is a standard cell ? Give an example.

Turn over

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**Ruhel Chisty MRACI CChem A**



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