



PRE-LAB QUESTIONS

1. What would you estimate the pH of hydrochloric acid, a strong acid, to be? What is the pH of water? The pH of HCl is 3.01, so yes a strong acid, the pH of water is 7
2. Use the Henderson-Hasselbalch equation to estimate the pH of a buffer solution that is composed of 20 mL of 10 M sodium formate and 20 mL of 1 M formic acid (Note: the pK_a of formic acid is 3.75). 4.75 would be the pH
3. Look at Table 4 in the procedure portion of the experiment. Calculate the pH you would expect each of the buffer solutions (A, B, C, D, and E) to be using the Henderson-Hasselbalch equation, assuming that the solutions of acetic acid and sodium acetate are equimolar.

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CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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