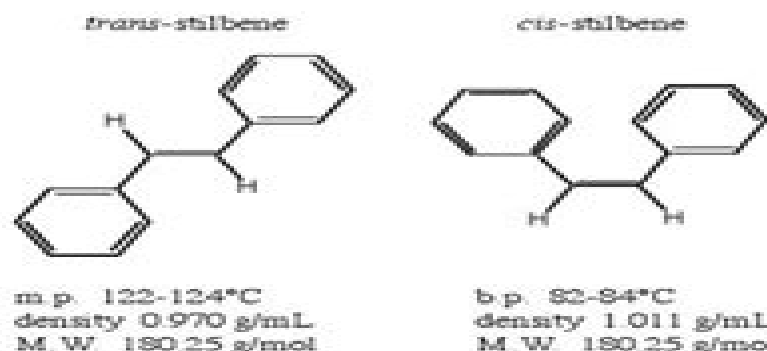


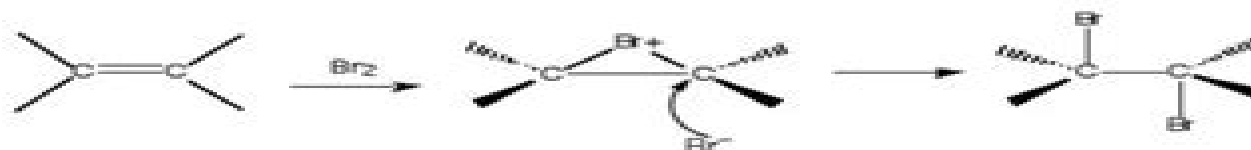
## Experiment 11: Bromination of Stilbene

In the experiment entitled Steam Distillation, you demonstrated that eugenol has one or more carbon-carbon double bonds by reacting it with bromine,  $\text{Br}_2$ , and observing the discharge of the reddish-brown color of bromine as each drop reacted with the double bond in eugenol. At that time, you were not asked to specify the structure of the product. Now that you have studied the mechanism of the addition of bromine to alkenes in lecture, as well as stereochemistry and optical isomerism, you are equipped to explore some of the more subtle aspects of this addition reaction.

In today's experiment, you will carry out the addition of bromine to the *trans* and *cis* isomers of 1,2-diphenylethene, more commonly known as *trans*- and *cis*-stilbene.



In both cases, the nucleophilic double bond undergoes an electrophilic addition reaction by the bromine reagent which proceeds via a cyclic bromonium ion. The addition of bromine begins at one side of the double bond (either side is equally likely, but only one option is drawn) and is followed by attack of bromide ion on the bromonium ion (again, attack could occur at either carbon since the ion is symmetric, but only one option is drawn). The result is a *trans* dibromide, as shown in the equation below.



Since the *cis* and *trans* isomers of stilbene have different geometries, it follows that upon reaction with bromine they give rise to stereoisomeric bromonium ions and, eventually, products that differ only by their stereochemistry.

The product of this reaction, 1,2-dibromo-1,2-diphenylethane, has two stereogenic centers which normally leads one to predict four ( $2^n = 2^2$  in this case) possible stereoisomers (i.e., four optical isomers). However, there are only three possible isomers in the case of 1,2-dibromo-1,2-diphenylethane since one, the so-called *meso* isomer, is

# Preparation Of Stilbene Dibromide Lab Report

**AW Rasmussen**



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**Experimental Organic Chemistry** John C. Gilbert, Stephen F. Martin, 2006 This proven and well tested laboratory manual for organic chemistry students contains procedures for both miniscale also known as small scale and microscale users This lab manual gives students all the necessary background to enter the laboratory with the knowledge to perform the experiments with confidence For the microscale labs experiments were chosen to provide tangible quantities of material which can then be analyzed Chapters 1 2 introduce students to the equipment record keeping and safety of the laboratory Chapters 3 6 and 8 are designed to introduce students to laboratory techniques needed to perform all experiments In Chapters 7 and 9 through 20 students are required to use the techniques to synthesize compounds and analyze their properties In Chapter 21 students are introduced to multi step syntheses of organic compounds a practice well known in chemical industry In Chapter 23 students are asked to solve structures of unknown compounds The new chapter 24 introduces a meaningful experiment into the textbook that reflects the increasing emphasis on bioorganic chemistry in the sophomore level organic lecture course This experiment not only gives students the opportunity to accomplish a mechanistically interesting and synthetically important coupling of two amino acids to produce a dipeptide but also provides valuable experience regarding the role of protecting groups in effecting synthetic transformations with multiple functionalized molecules

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