

EXPERIMENT 2 PRE-LAB QUESTIONS (20 points)

Stoichiometry and Thermochemistry

1. There are three possible unknowns for this experiment. Write balanced chemical equations for the reaction of each of these unknowns with sodium hypochlorite. (3 points)



2. If the stoichiometric ratio of your reaction were 1:1, what volumes of hypochlorite and your unknown would you use to get the highest temperature possible? What if it were 3:1? 4:1? (Hint: Remember the total volume will be 10 mL.) (3 points)

$$\text{NaClO} = \frac{1}{2} = 0.5 \times 10 \text{ mL} = 5 \text{ mL}$$

$$\text{unknown} = 5 \text{ mL}$$

$$\text{NaClO} = 0.25 \times 10 \text{ mL} = 2.5 \text{ mL}$$

$$\text{unk} = 0.75 \times 10 \text{ mL} = 7.5 \text{ mL}$$

$$\text{NaClO} = \frac{1}{3} \times 10 \text{ mL} = 3.3 \text{ mL}$$

$$\text{unk} = \frac{2}{3} \times 10 \text{ mL} = 6.7 \text{ mL}$$

3. If the stoichiometric ratio of your reaction were 1:1 what volumes would you use to collect four data points on either side of the highest temperature (stoichiometric ratio)? What if it were 3:1? 4:1? (3 points)

4. Finally, if you didn't know what the stoichiometric ratio of hypochlorite and your unknown was, what would be your approach? What combination of volumes could you use to determine the ratio between the hypochlorite ions and your unknown? (2 points)

Try different volumes until there is no leftover products or reactants.

5. If you combine 3 mL of water at 22°C and 7 mL of water at 24°C, what would the initial temperature of the solution be? (Hint: Find the weighted average.) (2 points)

$$3 \text{ mL} = 0.003 \text{ L @ } 295 \text{ K} \quad 7 \text{ mL} = 0.007 \text{ L @ } 297 \text{ K}$$

$$((0.003 \times 295) + (0.007 \times 297)) / 0.01 = 296.4 \text{ K}$$

$$\text{or } 23.4 \text{ °C}$$

6. When would you need to measure the temperature in each trial in order to calculate the change in temperature (ΔT)? (2 points)

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Richard Williams



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