

Naming Compounds Containing Polyatomic Ions

Use the following list of polyatomic ions to name the compounds given.

OH
 NO₃
 C₂H₃O₂

hydroxide
 nitrate
 acetate

ClO₃
 HCO₃
 SO₄
 CO₃

chlorate
 bicarbonate
 sulfate
 carbonate

SO₃
 PO₄
 NH₄

sulfite
 phosphate
 ammonium

- | | |
|--------------------------------------------------------|-----------------------------------------------------------|
| 1. CuSO ₄ _____ | 26. NaOH _____ |
| 2. KClO ₃ _____ | 27. NaNO ₃ _____ |
| 3. NH ₄ OH _____ | 28. Ba(OH) ₂ _____ |
| 4. K ₂ CO ₃ _____ | 29. K ₂ CO ₃ _____ |
| 5. Na ₂ SO ₄ _____ | 30. LiOH _____ |
| 6. KC ₂ H ₃ O ₂ _____ | 31. KHCO ₃ _____ |
| 7. NH ₄ Cl _____ | 32. NH ₄ F _____ |
| 8. Mg(OH) ₂ _____ | 33. KClO ₃ _____ |
| 9. Zn(NO ₃) ₂ _____ | 34. NaC ₂ H ₃ O ₂ _____ |
| 10. MgSO ₄ _____ | 35. CaClO ₃ _____ |
| 11. CaCO ₃ _____ | 36. SrNO ₃ _____ |
| 12. KOH _____ | 37. Be(OH) ₂ _____ |
| 13. Ca(OH) ₂ _____ | 38. Li ₂ SO ₄ _____ |
| 14. K ₂ CO ₃ _____ | 39. Al ₂ (SO ₄) ₃ _____ |
| 15. AlPO ₄ _____ | 40. BaSO ₃ _____ |
| 16. Na ₃ PO ₄ _____ | 41. NH ₄ I _____ |
| 17. MgSO ₄ _____ | 42. NH ₄ HCO ₃ _____ |
| 18. AgNO ₃ _____ | 43. Cu ₃ (PO ₄) ₂ _____ |
| 19. Na ₂ CO ₃ _____ | 44. Ni(OH) ₂ _____ |
| 20. NaHCO ₃ _____ | 45. ZnSO ₄ _____ |
| 21. Ca(NO ₃) ₂ _____ | 46. MnCO ₃ _____ |
| 22. BaSO ₄ _____ | 47. H ₂ SO ₄ _____ |
| 23. BPO ₃ _____ | 48. Au ₂ SO ₄ _____ |
| 24. LiNO ₃ _____ | 49. NH ₄ NO ₃ _____ |
| 25. Al(NO ₃) ₃ _____ | 50. LiClO ₃ _____ |

Polyatomic Ions Naming Answer

SJ Ball



Polyatomic Ions Naming Answer:

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Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career

Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make

available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields *General, Organic, and Biological Chemistry* Michael P. Garoutte, 2014-02-24

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