Name:

Neutralization Reactions Worksheet Breding: Creix 6 Chemistry III Lesson 4

Neutralization Reaction: Acid + Base -- Solt + Water H' from acid and OH from base combine to form H₂O. Lefterers form 'salt'. Remember in salt, write metal first: KCl not CIK!

Example 1: $HI + NaOH = H_2O + NaI$

H and Off combine to form water. One atom of I & one of Na leftower. Na is a metal so write NaI (not INa)

Example 2: __HF + __Ca(OH), --____ + _____

For complete neutralization, must have equal number of H and OH. Therefore we need 2 HF for 1 CatOHis.

_2_HF + _1_Cn(OH)₀ --____ + ____

Two H and two OH make two waters. Think of proof! Two boys and two girls make two cougles.

2 HF + 1 Ca(OH)2 - 2 H₂O +

What's left? 2 F atoms and 1 Ca atom form the 'sult'. Remorator that Ca is the metal, so it goes first.

 $2 \text{ HF} + 1 \text{ Ca(OH)}_2 = 2 \text{ HsO} + \text{CaF}_2$

Safety check: is CaF, a good formula? P.T. tells you lines are Ca⁻¹ and F, so read charge is 0 like it should be.

1. Complete & balance these neutralization reactions. Remember that H-O = HOH!

 $H_2SO_4 + Sr(OH)_5 \rightarrow \underline{\hspace{1cm}} H_2O + SrSO_4$

 $_Ca(OH)_2 + _H_0PO_4 - _H_2O + _Cay(PO_4)_2$

__HBr + Ba(OH); -- __HrO + ____

 $Zn(OH)_2 + \underline{\hspace{1cm}}HNO_5 \Rightarrow \underline{\hspace{1cm}} + Zn(NO_3)_2$

Al(OH)₃ + __HCl -- ____ + _____

2. Complete and balance the following equations representing neutralization reactions:

a) _CsOH + H₂SO₂ -- _H₂O + Cs₂CO₃

b) _HF + Mg(OH)₂ - _H₂O + ____

c) _HNO₂ + Al(OH)₃ - _H₂O + Al(NO₂)₃

d) HCl + KOH - + _____ +

e) HBrO₃ + LiOH - H₂O + _____

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Industrial Environmental Research Laboratory (Cincinnati, Ohio)

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