

Name: \_\_\_\_\_

Period: \_\_\_\_\_



1. Solute <b>A</b> more	A. The part that dissolves into the solution.
2. Solvent <b>B</b>	B. The part that dissolves something.
3. Covalent <b>C</b>	C. Compounds made up of two non-metals.
4. Ionic <b>D</b>	D. Since they are made up of ions, they are polar molecules.
5. Concentration <b>E</b> M m N	E. Tells you how much solvent is dissolved in the solution.

1. Water tension <b>F</b>	A. The breaking up of an ionic molecule by a polar solvent.
2. Dissociation <b>A</b>	B. A molecule that has a positive and negative end.
3. Hydrogen bonds <b>D</b>	C. The attractive force between two molecules of the same substance.
4. Polar Molecule <b>B</b> uneven	D. The attractive force caused by hydrogen protons to a more negative non-metal.
5. Capillary action <b>E</b>	E. What allows plants to draw water up from their roots.
6. Cohesion <b>C</b>	F. Caused by the cohesion of water molecules holding it together.

Give three ways to make sugar dissolve faster.

- stirring
- heating
- crushing

Which is the smallest particle size: rock salt, regular salt, or only powdered salt?

Which is the largest particle size: sugar cubes, granulated sugar (regular table sugar), or powdered sugar?

Which one will dissolve the fastest?

powdered sugar or granulated sugar?  
hot water or cold water?  
stirred or not stirred?  
 Large particles or small particles?

Which holds more solid solute: hot or cold liquids?  
 Which holds more gaseous solute: hot or cold solvents?

What property of gases allows them to be pressured into solution?  
can be compressed

1. Dissolve <b>E</b>	A. Causes more gaseous solvent to be forced into a liquid.
2. Pressure <b>A</b>	B. Increases solubility by moving the molecules around faster.
3. Crushing <b>C</b>	C. Increases solubility by allowing the solvent to touch more sides of the solute.
4. Heating <b>D</b>	D. Expands the solvent a bit to allow more room for the solute.
5. Stirring <b>B</b>	E. When something seems to disappear inside a solvent.

<u>N</u> Cooking oil	<u>P</u> Salt
<u>P</u> Soluble in water	<u>P</u> Ionic Compounds
<u>P</u> A compound with H's	<u>P</u> Dissolves in water
<u>P</u> Sugar <u>P/N</u>	<u>N</u> Insoluble in water

Draw a water molecule, being sure to label the ends as positive and negative.

$$\begin{array}{c} \text{H} \\ | \\ \text{O} \\ | \\ \text{H} \end{array}$$

Do metals become positive or negative?  
 Would a metal be attracted to water's hydrogen or oxygen?  
 Would a nonmetal be attracted to water's hydrogen or oxygen?

To which side of a Water Molecule are these Attracted?

<u>O</u> Magnesium	<u>H</u> Sulfur	<u>O</u> Iron
<u>H</u> Chlorine	<u>O</u> Potassium	<u>H</u> Bromine
<u>O</u> Calcium	<u>neither</u> Helium	<u>H</u> Nitrogen

Which is warmer: arctic water or tropical water?  
 Which holds more oxygen: arctic water or tropical water?

When power plants use cold streams for cooling, they can kill the fish by heating up the water. Why do many fish die? lack of oxygen (warm water causes the O<sub>2</sub> to be released in air)

Why are water bugs able to "walk on water"?  
water tension (cohesion)  
 How do plants get water from their roots up to their leaves?  
capillary action

# Night Packet Answer Key

**Wolfgang Guggemos**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name bar.

## Night Packet Answer Key:

Night Study Guide and Student Workbook BMI Staff,2010-09      **Curious Incident of the Dog in the Night: Study Guide and Student Workbook** ,2011-06-13      Twelfth Night Study Guide Saddleback Educational Publishing,2006-09-01

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