

Valence Electron Worksheet

Lewis Structures

Name _____

Date _____ Period _____

1. How many electrons are in the following elements?

A. Carbon

C. Indium

E. Tungsten

B. Chlorine

D. Calcium

F. Antimony

2. Draw Lewis Structures for each of the following elements. How many electrons are in their outer shells?

A. Silicon

C. Sulfur

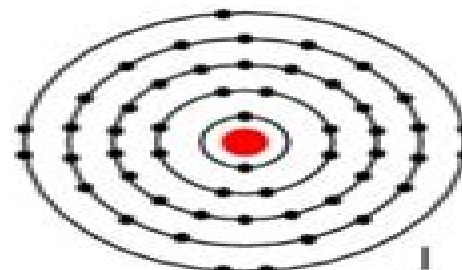
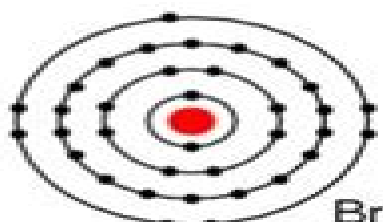
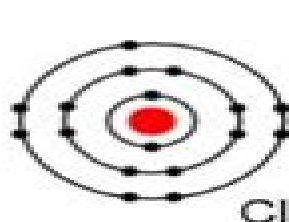
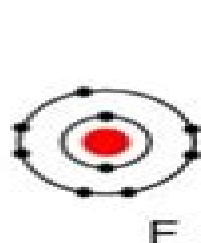
E. Beryllium

B. Potassium

D. Boron

F. Argon

3. Circle the valence electrons only. How many valence electrons does each atom have?



4. Finish the sentence: Elements in the same group have the same number of _____ electrons.

5. Fill in the electrons for the following elements. Use your work to answer how many valence electrons each element has.

Example: C (carbon): 1S $\uparrow\downarrow$ 2S $\uparrow\downarrow$ 2P \uparrow \uparrow \uparrow 3S _____ 3P _____ #valence electrons 4

He: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in helium _____

Cl: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in chlorine _____

Li: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in lithium _____

B: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in boron _____

F: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in fluorine _____

Ar: 1S _____ 2S _____ 2P _____ 3S _____ 3P _____ Number of valence electrons in argon _____

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VM Jensen



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Table of Contents Valence Electrons Worksheet Answers Chemistry If8766

1. Understanding the eBook Valence Electrons Worksheet Answers Chemistry If8766
 - The Rise of Digital Reading Valence Electrons Worksheet Answers Chemistry If8766
 - Advantages of eBooks Over Traditional Books
2. Identifying Valence Electrons Worksheet Answers Chemistry If8766
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Valence Electrons Worksheet Answers Chemistry If8766
 - User-Friendly Interface
4. Exploring eBook Recommendations from Valence Electrons Worksheet Answers Chemistry If8766

- Personalized Recommendations
 - Valence Electrons Worksheet Answers Chemistry If8766 User Reviews and Ratings
 - Valence Electrons Worksheet Answers Chemistry If8766 and Bestseller Lists
5. Accessing Valence Electrons Worksheet Answers Chemistry If8766 Free and Paid eBooks
 - Valence Electrons Worksheet Answers Chemistry If8766 Public Domain eBooks
 - Valence Electrons Worksheet Answers Chemistry If8766 eBook Subscription Services
 - Valence Electrons Worksheet Answers Chemistry If8766 Budget-Friendly Options
 6. Navigating Valence Electrons Worksheet Answers Chemistry If8766 eBook Formats
 - ePub, PDF, MOBI, and More
 - Valence Electrons Worksheet Answers Chemistry If8766 Compatibility with Devices
 - Valence Electrons Worksheet Answers Chemistry If8766 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Valence Electrons Worksheet Answers Chemistry If8766
 - Highlighting and Note-Taking Valence Electrons Worksheet Answers Chemistry If8766
 - Interactive Elements Valence Electrons Worksheet Answers Chemistry If8766
 8. Staying Engaged with Valence Electrons Worksheet Answers Chemistry If8766
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Valence Electrons Worksheet Answers Chemistry If8766
 9. Balancing eBooks and Physical Books Valence Electrons Worksheet Answers Chemistry If8766
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Valence Electrons Worksheet Answers Chemistry If8766
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Valence Electrons Worksheet Answers Chemistry If8766
 - Setting Reading Goals Valence Electrons Worksheet Answers Chemistry If8766
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Valence Electrons Worksheet Answers Chemistry If8766

- Fact-Checking eBook Content of Valence Electrons Worksheet Answers Chemistry If8766
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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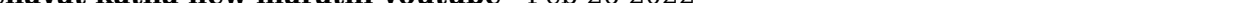
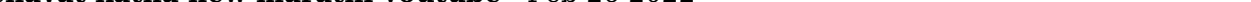
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