C11- PHYSICAL SCIENCE

Third Monthly Test Reviewer

Coverage: MODULE 1 UNIT 1: Atomic Theory UNIT 2: The Discovery and Synthesis of Chemical Elements

UNIT 1: Atomic Theory

 Atoms are composed of protons, neutrons, and electrons.

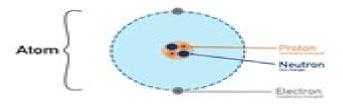


Table 2.3 Properties of Subatomic Particles

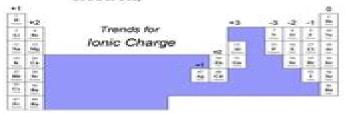
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- An atom is mostly empty space with a highly dense, positive center called nucleus.
- For a neutral atom, the number of protons is equivalent to the number of electrons.
 (Lithium has 3 protons therefore it also has 3 electrons)
- The atomic number of an element is also equivalent to its number of protons.
 (Nitrogen has an atomic number of 7 therefore it has 7 protons in its nucleus)
- The atomic mass (mass number) is primarily determined by the number of protons and neutrons.

- Isotopes are atoms with the same atomic number but different mass numbers. They differ in the number of neutrons.
 (Carbon-12, Carbon-13, Carbon-14)
- Ion is an element or atom that has a positive or negative net charge.
 It is charged because the number of electrons does not equal the number of protons in the atom.

Two types:

- Cation- positive ion (loses electron)
- Anion- negative ion (gains electron)



- Take note that when an atom is not electrically neutral, it means that it does not have the same number of protons and electrons.
- When an atom or element is positively charged, it means that it has more protons in its nucleus.

Historical Development of Atom

- Democritus & Leucippus: Atomism: All matter is made up of very small, indivisible particles called atoms (atomos).
- Aristotle: Four elements in nature (air, fire, water, and earth)
- John Dalton: Billiard Ball Model/Solid Sphere Model: Atomic Theory: All matter is made up of very small, indivisible particles called atoms.

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Brian Hopkins, Elena Geangu, Sally Linkenauger

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