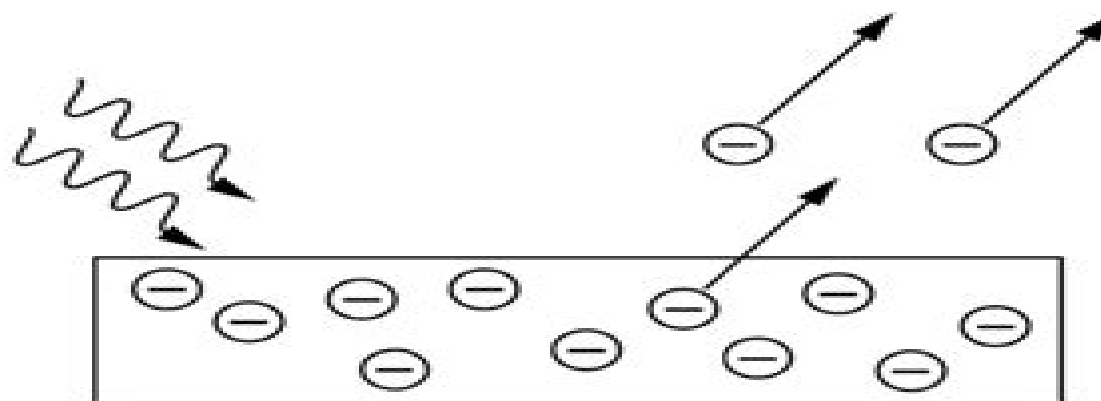


Photoelectric Effect Quiz



The diagram above represents the photoelectric effect for a metal. When the metal surface is exposed to light with increasing frequency and energy of photons, electrons first begin to be ejected from the metal when the energy of the photons is $3.3 \times 10^{-19} \text{ J}$.

1. Which of the following is closest to the frequency of the light with photon energy of $3.3 \times 10^{-19} \text{ J}$?

- (A) $5.0 \times 10^{-53} \text{ s}^{-1}$
 (B) $5.0 \times 10^{-16} \text{ s}^{-1}$
 (C) $5.0 \times 10^{14} \text{ s}^{-1}$
 (D) $5.0 \times 10^{52} \text{ s}^{-1}$

In photoelectric effect:

$$E = hf \quad \Rightarrow \quad f = \frac{E}{h} = \frac{3.3 \times 10^{-19} \text{ J}}{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}$$

$$f = 4.98 \times 10^{14} \text{ s}^{-1} = 5.0 \times 10^{14} \text{ s}^{-1}$$

Frequency of the light is $5.0 \times 10^{14} \text{ Hz}$

$$E = 3.3 \times 10^{-19} \text{ J}$$

$$h = \text{Planck constant} = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

- 2.

Color	Wavelength
Red	647 – 760 nm
Orange	585 – 647 nm
Yellow	575 – 585 nm
Green	491 – 575 nm
Blue	424 – 491 nm
Violet	300 – 424 nm

Using the wavelength information provided above, what is the color of the light?

$$E = 3.3 \times 10^{-19} \text{ J}$$

$$\lambda = 400 \text{ nm}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$c = 3 \times 10^8 \text{ m/s (velocity of light)}$$

$$E = hf = \frac{hc}{\lambda} \quad \Rightarrow \quad \lambda = \frac{hc}{E} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s} \times 3 \times 10^8 \text{ m/s}}{3.3 \times 10^{-19} \text{ J}}$$

$$\lambda = 6.0236 \times 10^{-7} \text{ m} = 6.0236 \times 10^2 \text{ nm} \left(\frac{10^9 \text{ nm}}{1 \text{ m}} \right) = 602.36 \text{ nm}$$

\Rightarrow Orange

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G Orfield



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Table of Contents Photoelectric Effect Quiz

1. Understanding the eBook Photoelectric Effect Quiz
 - The Rise of Digital Reading Photoelectric Effect Quiz
 - Advantages of eBooks Over Traditional Books
2. Identifying Photoelectric Effect Quiz
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Photoelectric Effect Quiz
 - User-Friendly Interface
4. Exploring eBook Recommendations from Photoelectric Effect Quiz
 - Personalized Recommendations
 - Photoelectric Effect Quiz User Reviews and Ratings
 - Photoelectric Effect Quiz and Bestseller Lists
5. Accessing Photoelectric Effect Quiz Free and Paid eBooks
 - Photoelectric Effect Quiz Public Domain eBooks
 - Photoelectric Effect Quiz eBook Subscription Services

- Photoelectric Effect Quiz Budget-Friendly Options
- 6. Navigating Photoelectric Effect Quiz eBook Formats
 - ePub, PDF, MOBI, and More
 - Photoelectric Effect Quiz Compatibility with Devices
 - Photoelectric Effect Quiz Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Photoelectric Effect Quiz
 - Highlighting and Note-Taking Photoelectric Effect Quiz
 - Interactive Elements Photoelectric Effect Quiz
- 8. Staying Engaged with Photoelectric Effect Quiz
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Photoelectric Effect Quiz
- 9. Balancing eBooks and Physical Books Photoelectric Effect Quiz
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Photoelectric Effect Quiz
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Photoelectric Effect Quiz
 - Setting Reading Goals Photoelectric Effect Quiz
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Photoelectric Effect Quiz
 - Fact-Checking eBook Content of Photoelectric Effect Quiz
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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