

Periodic Trends Practice



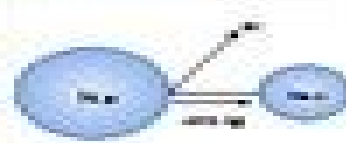
Atomic Radius

The distance from the nucleus of one bonded atom.



Electronegativity

How much an atom "wants" electrons or the atom's tendency to attract electrons.



Ionization Energy

The amount of energy required to remove electrons from an atom.



Electron Affinity

The amount of energy released when a neutral gaseous atom gains an electron.



Driving Question

How can the periodic table be used to predict atomic characteristics such as atomic radius, electronegativity, ionization energy, and electron affinity?

Describe what causes this trend...

As you move to the right, the number of protons increases, which increases the positive charge of the nucleus. This increases the attraction for electrons, which decreases the atomic radius.

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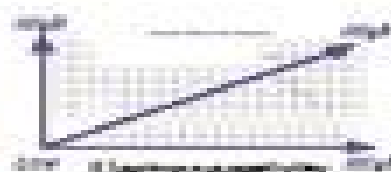
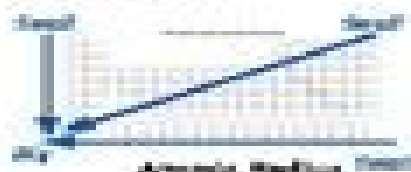
As you move to the right, the number of protons increases, which increases the positive charge of the nucleus. This increases the attraction for electrons, which decreases the atomic radius.

Shielding?

What is shielding? What effect does it have on the periodic trends defined above?

Shielding occurs when inner electrons block out the positive pull of the nucleus, making it harder for the outer electrons to feel the attraction. This causes the atomic radius to increase, electronegativity to decrease, ionization energy to decrease, and electron affinity to also decrease.

Add arrows to each of the following periodic tables to indicate the direction of the trend as it increases or decreases from left to right and top to bottom. (For Fig. 1 use a different color for each of the periodic tables so you can easily remember them.)



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Michael Brown



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