Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1.Calculate to required to mol Al in to 1.202(g)	2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO₃). 2KClO₃ → 2KCl + 3O₂								
14.8 moles Al 3 moles of O ₂ = 11.1			L1.1 mol O₂	0.61 mo	les KClO,	3 mo	oles of O2	_	
	4 more At				2 mole KClO ₃ 0.915 mol O ₂				
hydrochlo	ric acid ide (Zn¢ 2HCl _(aq)	, how many II ₂) are prod	duced?	must iodide chlori	be react (NaI) if de (NaCI +Cl ₂	ed w 10.0) are		s sodium f sodium	
20 g Zn 1 n	mail 20m 1	1 mole ZnCl ₂	136 g ZnO ₀	1					
4654	p. 20m 1	L rmusi Zin	1 moi ZnCl ₂	10 g NaCi	1 mg/ Nacl	2.0	mail Clu	71 g Clo	
= 41.8 g ZnCl ₂					S8-Sg NoCl	2 -	mol NaCi	1 mei Cl	
					64	07.p	Cly		
to replace 4 which are di	.00 grad	ms of silver		chieri	neded to ne gas to de gas?	reac e pro	of hydron t with 15 duce hyd	1.1.0	
	1 must AgNO:	1 male Cu	63.5 g Cu	15.1 9 <			1 mole H		
	170g AgNOr	2 mai AgNO:	1 mei Cu						
					-	0.2	1 mol	es.H ₂	

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Kenneth W. Raymond

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