

Specific Heat Problems – ANSWERS

The last page of this document has Miscellaneous Information that you will need for these problems.

1. How many calories are needed to warm 55.8 g of water at 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. H. \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g}^\circ\text{C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g}^\circ\text{C} \times 52.1^\circ\text{C}$$

$$\text{Heat} = 2907.18 \text{ cal} \quad \text{convert to scientific notation}$$

$$= 2.9\textbf{0}718 \times 10^3 \text{ cal}$$

$$= 2.91 \times 10^3 \text{ cal (3 sig. figs.)}$$

2. How many calories are needed to warm 55.8 g of alcohol from 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. H. \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g}^\circ\text{C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g}^\circ\text{C} \times 52.1^\circ\text{C}$$

$$\text{Heat} = 1686.16 \text{ cal} \quad \text{convert to scientific notation}$$

$$\text{Heat} = 1.6\textbf{8}616 \times 10^3 \text{ cal}$$

$$\text{Heat} = 1.69 \times 10^3 \text{ cal}$$

Practice Specific Heat Problems With Answers

John Bird



Practice Specific Heat Problems With Answers:

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Municipal Engineering: Sanitary Record and the Municipal Motor , 1928 **Math Concepts for Food Engineering** Richard W. Hartel, D.B. Hyslop, T.A. Howell

Jr., 2008-03-17 A Supplement for Food Science Engineering Students Who Need to Improve Their Mathematical Skills A remedial textbook for understanding mathematical theories and formulas Math Concepts for Food Engineering Second Edition helps students improve their mathematical skills so that they can succeed in food engineering cour *A New Approach to I.C.S.E. Physics for Class X* V.K. Sally, R.N. Das Gupta, Goyal Brothers Prakashan, 2019-12-11 Goyal Brothers Prakashan The Pearson Guide to Objective Physics for the IIT-JEE 2012: (All India Engineering Entrance Examination) ,

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Table of Contents Practice Specific Heat Problems With Answers

1. Understanding the eBook Practice Specific Heat Problems With Answers
 - The Rise of Digital Reading Practice Specific Heat Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Specific Heat Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Specific Heat Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Specific Heat Problems With Answers

- Personalized Recommendations
- Practice Specific Heat Problems With Answers User Reviews and Ratings
- Practice Specific Heat Problems With Answers and Bestseller Lists
- 5. Accessing Practice Specific Heat Problems With Answers Free and Paid eBooks
 - Practice Specific Heat Problems With Answers Public Domain eBooks
 - Practice Specific Heat Problems With Answers eBook Subscription Services
 - Practice Specific Heat Problems With Answers Budget-Friendly Options
- 6. Navigating Practice Specific Heat Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Specific Heat Problems With Answers Compatibility with Devices
 - Practice Specific Heat Problems With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Specific Heat Problems With Answers
 - Highlighting and Note-Taking Practice Specific Heat Problems With Answers
 - Interactive Elements Practice Specific Heat Problems With Answers
- 8. Staying Engaged with Practice Specific Heat Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Specific Heat Problems With Answers
- 9. Balancing eBooks and Physical Books Practice Specific Heat Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Specific Heat Problems With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Specific Heat Problems With Answers
 - Setting Reading Goals Practice Specific Heat Problems With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Specific Heat Problems With Answers

- Fact-Checking eBook Content of Practice Specific Heat Problems With Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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