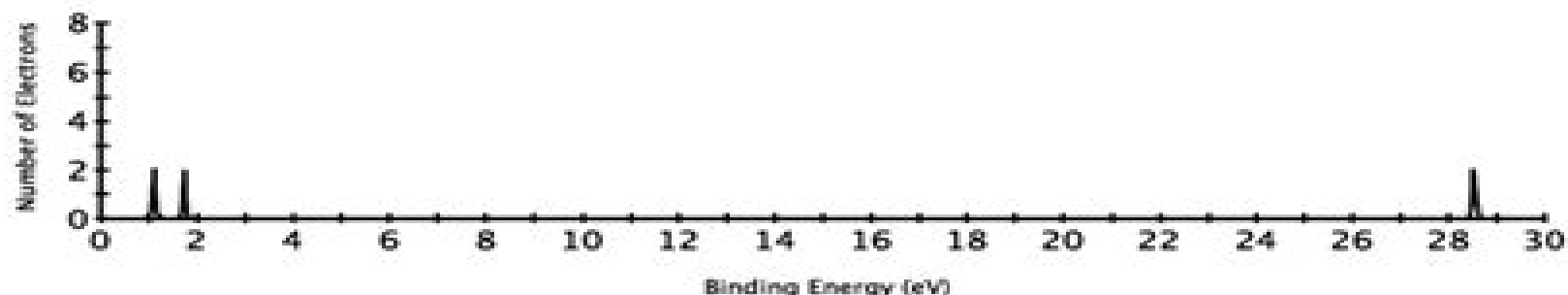


Question:

Below is the photoelectron spectrum of carbon:



Which of the following is the best explanation for the large difference in binding energy between the peaks on the left and the right.

- a. The electrons generating the peak on the right are further away from the nucleus.
- b. There are four electrons all together in the group on the left.
- c. The electrons generating the peak on the right have greater electron shielding.
- d. The electrons generating the peak on the left are further from the nucleus and experience greater electron shielding.

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Answer:

The correct answer is “d”, the electrons generating the peak on the left are further from the nucleus and experience greater electron shielding. The peaks on the left are generated by the ionization of electrons in the 2s and 2p orbitals. These electrons are further from the nucleus and are shielded from the attractive force of the nucleus by the electrons in the 1s orbital. The greater distance and shielding reduce the coulombic attraction making it easier to remove the electrons in the 2s and 2p orbitals from the atom.

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