

## Photoelectron Spectroscopy (PES) ANSWER KEY

1. In a photoelectron spectrum, photons of 165.7 MJ/mol strike atoms of an unknown element. If the kinetic energy of the ejected electrons is 25.4 MJ/mol, what is the ionization energy of the element?

$$\text{IE} = \text{Energy of Photons} - \text{KE of electrons} = 165.7 \text{ MJ/mol} - 25.4 \text{ MJ/mol} = 140.3 \text{ MJ/mol}$$

2. What determines the position and the height (intensity) of each peak in a photoelectron spectrum?

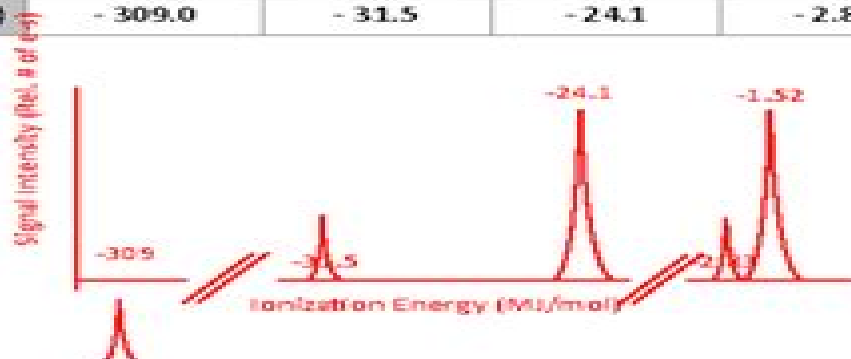
Position = Ionization Energy, height = relative number of electrons

3. Why is the distance of the energy level from the nucleus important in determining the corresponding peak position in the photoelectron spectrum?

When electrons are further from the nucleus, the interactions (Coulombic attractions) between the electrons and nucleus are weaker, so less energy is needed to remove them.

4. Based on the information provided below, draw a photoelectron spectrum for argon. Indicate the relative intensities and positions of all peaks.

e- conf	1s <sup>2</sup>	2s <sup>2</sup>	2p <sup>6</sup>	3s <sup>2</sup>	3p <sup>6</sup>
IE (MJ/mol)	-309.0	-31.5	-24.1	-2.83	-1.52



5. Identify the element which produced the photoelectron spectrum shown below. Briefly explain your reasoning. Label each peak with the portion of the electron configuration that it represents.

# Photoelectron Spectroscopy Activity With Answer Key

**G Orfield**



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