

### Balancing Equations Worksheet

- 1)  $\text{___ C}_3\text{H}_8 + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- 2)  $\text{___ Al} + \text{___ Fe}_3\text{N}_2 \rightarrow \text{___ AlN} + \text{___ Fe}$
- 3)  $\text{___ Na} + \text{___ Cl}_2 \rightarrow \text{___ NaCl}$
- 4)  $\text{___ H}_2\text{O}_2 \rightarrow \text{___ H}_2\text{O} + \text{___ O}_2$
- 5)  $\text{___ C}_6\text{H}_{12}\text{O}_6 + \text{___ O}_2 \rightarrow \text{___ H}_2\text{O} + \text{___ CO}_2$
- 6)  $\text{___ H}_2\text{O} + \text{___ CO}_2 \rightarrow \text{___ C}_7\text{H}_8 + \text{___ O}_2$
- 7)  $\text{___ NaClO}_3 \rightarrow \text{___ NaCl} + \text{___ O}_2$
- 8)  $\text{___ (NH}_4\text{)}_3\text{PO}_4 + \text{___ Pb(NO}_3\text{)}_4 \rightarrow \text{___ Pb}_3\text{(PO}_4\text{)}_4 + \text{___ NH}_4\text{NO}_3$
- 9)  $\text{___ BF}_3 + \text{___ Li}_2\text{SO}_3 \rightarrow \text{___ B}_2\text{(SO}_3\text{)}_3 + \text{___ LiF}$
- 10)  $\text{___ C}_7\text{H}_{17} + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- 11)  $\text{___ CaCO}_3 + \text{___ H}_3\text{PO}_4 \rightarrow \text{___ Ca}_3\text{(PO}_4\text{)}_2 + \text{___ H}_2\text{CO}_3$
- 12)  $\text{___ Ag}_2\text{S} \rightarrow \text{___ Ag} + \text{___ S}_8$
- 13)  $\text{___ KBr} + \text{___ Fe(OH)}_3 \rightarrow \text{___ KOH} + \text{___ FeBr}_3$
- 14)  $\text{___ KNO}_3 + \text{___ H}_2\text{CO}_3 \rightarrow \text{___ K}_2\text{CO}_3 + \text{___ HNO}_3$
- 15)  $\text{___ Pb(OH)}_4 + \text{___ Cu}_2\text{O} \rightarrow \text{___ PbO}_2 + \text{___ CuOH}$
- 16)  $\text{___ Cr(NO}_2\text{)}_2 + \text{___ (NH}_4\text{)}_2\text{SO}_4 \rightarrow \text{___ CrSO}_4 + \text{___ NH}_4\text{NO}_2$
- 17)  $\text{___ KOH} + \text{___ Co}_3\text{(PO}_4\text{)}_2 \rightarrow \text{___ K}_3\text{PO}_4 + \text{___ Co(OH)}_2$
- 18)  $\text{___ Sn(NO}_2\text{)}_4 + \text{___ Pt}_3\text{N}_4 \rightarrow \text{___ Sn}_3\text{N}_4 + \text{___ Pt(NO}_2\text{)}_4$
- 19)  $\text{___ B}_2\text{Br}_6 + \text{___ HNO}_3 \rightarrow \text{___ B(NO}_3\text{)}_3 + \text{___ HBr}$
- 20)  $\text{___ ZnS} + \text{___ AlP} \rightarrow \text{___ Zn}_3\text{P}_2 + \text{___ Al}_2\text{S}_3$

# Practice Balancing Chemical Equations

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everyday substances Canny I've heard of acids and bases They have something to do with the pH scale right Jammy Exactly The pH scale measures the acidity or basicity of a substance It ranges from 0 to 14 with 7 being neutral Substances with a pH less than 7 are acidic while those with a pH greater than 7 are basic Canny I'm starting to see how chemistry is all around us shaping the world we live in Jammy Absolutely Chemistry influences everything from the food we eat the medicines we take the materials we use and even the air we breathe It's a central science that connects many other fields of study Canny This is truly captivating I can't wait to learn more about the wonders of chemistry and delve into its intricacies Jammy I'm thrilled that you're so enthusiastic Canny In the upcoming chapters we'll explore topics like thermodynamics organic chemistry chemical kinetics and even the cutting edge world of quantum chemistry Canny I'm ready for the adventure Let's uncover the mysteries of the fascinating world of chemistry together Jammy Jammy You bet Get ready for an amazing journey into the heart of matter and the magic of molecules Chemistry awaits



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