	9.	 At 700 °C, the equilibrium constant (K_c) for the reaction N_c (g) + 3 H_c (g) ≠ 2 NH_c (g) is 2.37 × 10°. If i mol of N_c, 3.20 mol of H_c and 0.632 mole of NH_c are mixed in a 2.00 L container at 700 °C: 		
		a. the value of Q is unknown.d. the reaction is at equilibrium.		c. the value of Q is equal to K _c .

- _____ 10. The only factor that causes a change in the numerical value of an equilibrium constant is a(n): a. change in volume. b. change in pressure. c. change in concentrations. d. addition of a catalyst. e. change in temperature.
- 11. Which of the items listed below would increase the yield of NO₂ in the following equilibrium reaction?

$$2 \text{ NO (g)} + O_2 (g) \rightleftharpoons 2 \text{ NO_2 (g)}$$

(exothermic)

- Add more O₂ to the equilibrium mixture.
- Warm the reaction mixture.
- III. Increase the pressure by decreasing the volume of the reaction vessel.

a. II only b. I and II c. I and III d. II and III e. All of these.

- 12. For the reaction H₂ (g) + CO₂ (g) ₹² H₂O (g) + CO (g) at 700 °C, K₂ = 0.534. Calculate the concentration of H₂ at equilibrium if a mixture of 0.200 mol of H₂O, 0.200 mol CO, 0.750 mol H₂ and 0.750 mol CO₂ is heated to 700 °C in a 1.00 liter container.
 a. 0.500 M b. 0.100 M c. 0.573 M d. 0.132 M e. 0.735 M
- The conjugate acid of HPO₄² is:
 H₂PO₄ b. H₂PO₄ c. PO₄³ d. H⁴ e. none of these
 - 14. If the reaction below proceeds, which of the following is true?

 $NH_3 + C_2H_3OH \rightarrow NH_3^- + C_2H_3OH_3^-$

a. C_2H_3OH is an acid. b. NH_2 is the conjugate base of NH_3 . c. NH_3 is a base. d. $C_2H_3OH_2$ is a base. e. none of these

- ____ 15. Which aqueous solution(s) below is(are) considered basic?
 - I. A solution with a pH = 6.
 - II. A solution with [OH] = 1 × 10⁻⁷
 - III. A solution with $[H_3O^*] = 1 \times 10^{-9}$

a. I only b. II only c. III only d. I and II e. I and III

16. The pH of a diet Coca Cola solution is 2.72. What is the concentration of [OH] present in this solution?
a. [OH] = 5.25 × 10⁴² M
b. [OH] = 1.91 × 10²³ M
c. [OH] = 1.3 × 10⁴³ M
d. [OH] = 7.6 × 10⁴⁴ M
e. [OH] = 3.16 × 10³⁶ M

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Zak Bagans

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