# 11-2 Practice Problems

 Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:

Eglin x 1 mai Liju x 3 mai Liot x 24 g Liot = 0.78 g

 What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?

029 NoI + Cl2 - 2NoCI + Iz 029 NoI + Cl2 - 2NoCI + Iz 029 NoI x 1mol x 2mol NoI (mol NoC) NoCI

> Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:

0+85gCyHo x 1mol x 5mol Cox x 44g = 2.58g 1 x 58g x 2mol CyHo 1mol CO z

 Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:

0.4695bD= x [mol x 2mol 5b, 03 x 121.769 fa.389 Sb

5. What mass of hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>)
must decompose to produce 0.77 g of
water?
2H<sub>2</sub>O<sub>2</sub> -> 2H<sub>2</sub>O + O<sub>2</sub>

0.779 H20 x (mol x 2420 x 349 = [.4594:02

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?

0.69gCOz × 1mol × 2mol CO x 289 = 0.449 1 449 2mol COz 1mol = 0.449  Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:

 $\frac{N8(NO_3)_2 + 2N8OH \rightarrow N8(OH)_2 + 2N8NO_3}{O.73_3 V80_3)_2 \times \frac{1000}{185_3} \times \frac{2001 N8NO_3}{1001 N6NO_3)_2} \times \frac{85_3}{1001} = \frac{(2.68_3)}{1001 N6NO_3}$ 

 Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:

 $0.64_9H_2O \times \frac{CaC_2 + 2H_2O - Ca(OH)_2 + C_2H_2}{189} \times \frac{1001(a(0H)_2 + C_2H_2}{1001} \times \frac{749}{1001} = 1.329$   $Ca(OH)_2$ 

How many grams of ozone (O<sub>3</sub>) must decompose to produce 0.87 g of oxygen?

0.8790 x 1mg/ x 2mg/ x 48903 = 0.87909

 Find the mass of sugar (C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation: C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> → 2C<sub>2</sub>H<sub>6</sub>O + 2CO<sub>2</sub>

 How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP?
 The balanced equation is S + O<sub>2</sub> → SO<sub>2</sub>.

 Find the mass of benzene (C<sub>6</sub>H<sub>6</sub>) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:

2C4H6 + 15O2 → 6H2O + 12CO2

# **Practice Problem 30 Chapter 12 Chemistry Prentice Hall**

**Clemens Wendtner** 

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Fundamentals of Environmental Sampling and Analysis Chunlong Zhang,2024-04-02 Fundamentals of Environmental Sampling and Analysis A fully reworked and updated introduction to the fundamentals and applications of environmental sampling and analysis Environmental sampling and analysis are essential components of environmental data acquisition and scientific research The acquisition of reliable data with respect to proper sampling chemical and instrumental methodology and QA QC is a critical precursor to all environmental work No would be environmental scientist engineer or policymaker can succeed without an understanding of how to correctly acquire assess and use credible data Fundamentals of Environmental Sampling and Analysis 2nd edition provides this understanding with a comprehensive survey of the theory and applications of these critical sampling and analytical tools The field of environmental research has expanded greatly since the publication of the first edition and this book has been completely rewritten to reflect the latest studies and technological developments The resulting mix of theory and practice will continue to serve as the standard introduction to the subject

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Intermediate Algebra Linda Gilbert, Jimmie Gilbert, 1987 The Publishers Weekly .1919 Contact Angle, **Wettability and Adhesion** K. L. Mittal, 2009-09-30 The topic of wettability measured in terms of contact angle is of tremendous interest from both fundamental and applied points of view Wettability plays an essential role in many industrial processes so an understanding of factors dictating wettability and how to modulate it is of paramount importance In the last years there has been an explosive interest in superhydrophobic surfaces i e surfaces with water contact angle of 150 or higher because of their relevance importance in many areas ranging from self cleaning windows to nanofluidics Also recently there has been heightened activity in the field of electrowetting Contact Angle Wettability and Adhesion Volume 6 is divided into four parts Part 1 Fundamental Aspects Part 2 Wettability Control Modification Part 3 Superhydrophobic Surfaces and Part 4 Surface Free Energy and Relevance of Wettability in Adhesion The topics covered include a guide to the equilibrium contact angles maze fundamental aspects of wetting of rough and chemically heterogeneous surfaces work of adhesion for rock oil brine systems Is the world basic wettability control modification using various approaches superhydrophobic surfaces and ways to impart superhydrophobicity adsorption on superhydrophobic surfaces solid surface energy determination surface modification of different materials relevance of wettability and adhesion aspects in a variety of reinforced composites In essence this volume reflects the cumulative wisdom of many active and renowned researchers and provides a commentary on contemporary research in the fascinating world of contact angles and wettability This volume and its predecessors 5 volumes containing bountiful information will be of much value to anyone interested involved in controlling wetting phenomena and their applications Catalog of Copyright Entries. Third Series Library of Congress. Copyright Office, 1967 Includes Part 1 Number 1 Books and Pamphlets Including Serials and Contributions to Periodicals January June

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