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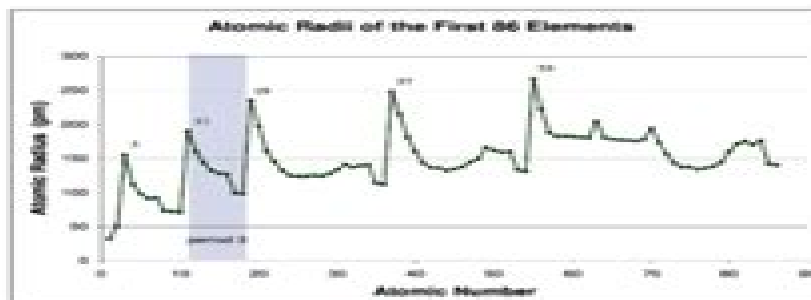
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### trends on the periodic table – atomic radius

The **atomic radius** describes the distance between the nucleus and the outermost electron. This distance is very small so it is often reported in **picometers** ( $1 \times 10^{-12}$  m) or **angstroms** ( $1 \times 10^{-10}$  m). One way of measuring the atomic radius is to observe the distance between two bonding atoms. The distance between the nuclei of the bonding atoms is equal to twice the atomic radius. This distance is called the **bonding atomic radius** and is shorter than the non-bonded radius.



A pattern is observed when the atomic radius is plotted versus the atomic number, as seen above. As might be expected the largest atoms are found at the bottom of the periodic table, however, as you go across a period, the atomic radius decreases (see highlighted region in the graph above). All of the elements in a period have the same orbitals which have the same energy and same size. Additionally, the elements that lay on the right side of the periodic table have nuclei which contain more protons. This results in a larger attraction of the electrons and a smaller radius. This explains why argon (atomic number = 18) has a smaller radius than sodium (atomic number = 11).

**Directions:** Answer the following questions on a separate piece of paper:

1. Estimate the size of the smallest atom in the chart above. Estimate the size of the largest.
2. Which five elements are represented by the tallest 'peaks' in the graph above?
3. Write the approximate radius for each of the noble gases in the table below:

element	atomic radius	element	atomic radius	element	atomic radius
helium		argon		xenon	
neon		krypton		radon	

4. What is the general trend for atomic radius as you go down the noble gas family?
5. What is the general trend as you move through period 3 (from element 11 to element 18)?

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**Rob Ritchie**



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## **Table of Contents Practice Problems With Periodic Trends**

1. Understanding the eBook Practice Problems With Periodic Trends
  - The Rise of Digital Reading Practice Problems With Periodic Trends
  - Advantages of eBooks Over Traditional Books
2. Identifying Practice Problems With Periodic Trends
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Practice Problems With Periodic Trends
  - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Problems With Periodic Trends
  - Personalized Recommendations
  - Practice Problems With Periodic Trends User Reviews and Ratings
  - Practice Problems With Periodic Trends and Bestseller Lists

5. Accessing Practice Problems With Periodic Trends Free and Paid eBooks
  - Practice Problems With Periodic Trends Public Domain eBooks
  - Practice Problems With Periodic Trends eBook Subscription Services
  - Practice Problems With Periodic Trends Budget-Friendly Options
6. Navigating Practice Problems With Periodic Trends eBook Formats
  - ePub, PDF, MOBI, and More
  - Practice Problems With Periodic Trends Compatibility with Devices
  - Practice Problems With Periodic Trends Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Practice Problems With Periodic Trends
  - Highlighting and Note-Taking Practice Problems With Periodic Trends
  - Interactive Elements Practice Problems With Periodic Trends
8. Staying Engaged with Practice Problems With Periodic Trends
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Practice Problems With Periodic Trends
9. Balancing eBooks and Physical Books Practice Problems With Periodic Trends
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Practice Problems With Periodic Trends
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Practice Problems With Periodic Trends
  - Setting Reading Goals Practice Problems With Periodic Trends
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Practice Problems With Periodic Trends
  - Fact-Checking eBook Content of Practice Problems With Periodic Trends
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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