

Practice Problem!!

1. If 12.5 g of $\text{Cu}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ is added to 500 mL of 1.00 M aqueous ammonia, to form $[\text{Cu}(\text{NH}_3)_4]^{+2}$. The stability constant, K_f for $[\text{Cu}(\text{NH}_3)_4]^{+2}$ is 2.1×10^{13} what is the equilibrium concentration of $\text{Cu}^{2+}(\text{aq})$ ion in the solution?



Answer: $[\text{Cu}^{+2}] = 2.1 \times 10^{-14} \text{ M}$

2. The ferrocyanide ion $[\text{Fe}(\text{CN})_6]^{4-}$ is very stable, with a K_f of 1×10^{35} . Calculate the concentration of cyanide ion in equilibrium with a 0.65 M solution of $\text{K}_4[\text{Fe}(\text{CN})_6]$.



Answer: $2 \times 10^{-6} \text{ M}$

Practice Solubility Problems

**Barron's Educational Series, Mark
Kernion, Joseph A. Mascetta**



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