# 4

### Atomic Structure

Reviewsheet

#### A. Completion

Use this completion exercise to check your understooding of the concepts and terms that are introduced in this chapter. Each black can be completed with a term, shore phoses, or number.

It was \_\_\_7\_\_ who discovered the nucleus of the atom. The nucleus has a \_\_\_8\_\_ charge and it occupies a very small volume of the atom. In contrast, the negatively charged \_\_\_9\_\_ occupy most of the volume of the atom.

The number of \_\_19\_ in the nucleus of the atom is the atomic \_\_11\_ of that element: Because atoms are electrically neutral, the number of protons and \_\_12\_ in an atom are equal. The sum of the \_\_13\_ and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of \_\_14\_ in the nucleus. Atoms that have the same number of protons but different mass numbers are called \_\_15\_\_.

The \_\_\_16\_\_ of an element is the weighted average of the \_\_2 masses of the isotopes of that element. Two isotopes of sulfur are \_\_12 S and \_\_14 S. An atom of the sulfur—32 isotope contains \_\_17\_\_ protons and \_\_18\_\_ neutrons. The sulfur—34 isotope has \_\_19\_\_ protons and \_\_20\_\_ neutrons.

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4.	-cxotoos	4.2
	736174COLLE	4.2
6.	nucleus +	4.3
7.	Rutnerford	4.3
8.	positive	4.3
9.	electrons	4.8
10.	prodores (elec)	4.4
11.	rumber	4.4
12.	exchana	4.4
13.	protons Celec	35
14.	neutrons	4.6
15.	isotopes	4.6
16. 3	atomic mass.	4.7
(7	16	4.5
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**Lingjun Ying** 

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