

**CHAPTER 12 REVIEW***Solutions***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Solid  $\text{CaCl}_2$  does not conduct electricity. Explain why it is considered to be an electrolyte.

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2. Explain the following statements at the molecular level:

a. Generally, a polar liquid and a nonpolar liquid are immiscible.

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b. Carbonated soft drinks taste flat when they warm up.

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3. An unknown compound is observed to mix with toluene,  $\text{C}_6\text{H}_5\text{CH}_3$ , but not with water.

a. Is the unknown compound ionic, polar covalent, or nonpolar covalent? Explain your answer.

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b. Suppose the unknown compound is also a liquid. Will it be able to dissolve table salt? Explain why or why not.

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# Review Solutions Mixed Review Short Answers

**Marcel A. Müller**



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