Balancing Equations and Simple Stoichiometry-KEY

Balance the following equations:

- 1 N₂ + 3 F₂ → 2 NF₃
- 2) 2 C₆H₁₀ + 17 O₂ → 12 CO₂ + 10 H₂O
- 3) 1 HBr + 1 KHCO₃ → 1 H₂O + 1 KBr + 1 CO₂
- 4) 2 GaBr₃ + 3 Na₂SO₃ → 1 Ga₂(SO₃)₃ + 6 NaBr
- 5) 3 SnO + 2 NF₃ → 3 SnF₂ + 1 N₂O₃

Using the following equation:

- 6) How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?
- ? g Na₂SO₄ = 200 g NaOH→mol NaOH→mol Na₂SO₄ → g Na₂SO₄

Using the following equation:

$$Pb(SO_4)_2 + 4 LiNO_3 \rightarrow Pb(NO_3)_4 + 2 Li_2SO_4$$

How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

Molar masses: LiNO₃ = 68.95 g/mol & Li₂SO₄ = 109.9 g/mol ? g LiNO₃ = 250 g Li₂SO₄ \rightarrow mol Li₂SO₄ \rightarrow mol LiNO₃ \rightarrow g LiNO₃

$$250 \text{ g Li}_2SO_4 \times \underline{\text{I mol Li}_2SO_4} = 2.27 \text{ mol Li}_2SO_4 \times \underline{4 \text{ mol LiNO}_3} = 4.54 \text{ mol LiNO}_3$$

 $109.9 \text{ g Li}_2SO_4 = 2.27 \text{ mol Li}_2SO_4 \times \underline{4 \text{ mol LiNO}_3} = 4.54 \text{ mol LiNO}_3$

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- 1. Understanding the eBook Review Sheet Stoichiometry Answers Modern Chemistry
 - The Rise of Digital Reading Review Sheet Stoichiometry Answers Modern Chemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Review Sheet Stoichiometry Answers Modern Chemistry
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Sheet Stoichiometry Answers Modern Chemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Review Sheet Stoichiometry Answers Modern Chemistry
 - Personalized Recommendations
 - Review Sheet Stoichiometry Answers Modern Chemistry User Reviews and Ratings

- Review Sheet Stoichiometry Answers Modern Chemistry and Bestseller Lists
- 5. Accessing Review Sheet Stoichiometry Answers Modern Chemistry Free and Paid eBooks
 - Review Sheet Stoichiometry Answers Modern Chemistry Public Domain eBooks
 - Review Sheet Stoichiometry Answers Modern Chemistry eBook Subscription Services
 - Review Sheet Stoichiometry Answers Modern Chemistry Budget-Friendly Options
- 6. Navigating Review Sheet Stoichiometry Answers Modern Chemistry eBook Formats
 - o ePub, PDF, MOBI, and More
 - Review Sheet Stoichiometry Answers Modern Chemistry Compatibility with Devices
 - Review Sheet Stoichiometry Answers Modern Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Sheet Stoichiometry Answers Modern Chemistry
 - Highlighting and Note-Taking Review Sheet Stoichiometry Answers Modern Chemistry
 - Interactive Elements Review Sheet Stoichiometry Answers Modern Chemistry
- 8. Staying Engaged with Review Sheet Stoichiometry Answers Modern Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Sheet Stoichiometry Answers Modern Chemistry
- 9. Balancing eBooks and Physical Books Review Sheet Stoichiometry Answers Modern Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Sheet Stoichiometry Answers Modern Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Sheet Stoichiometry Answers Modern Chemistry
 - Setting Reading Goals Review Sheet Stoichiometry Answers Modern Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Sheet Stoichiometry Answers Modern Chemistry
 - Fact-Checking eBook Content of Review Sheet Stoichiometry Answers Modern Chemistry
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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