

Periodic Trends

1. Explain the following trends:

- Atomic Radius
- Electronegativity
- Ionization Energy

2. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

3. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

4. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

5. Order the following elements by increasing atomic radius:

Be, Mg, Ca, Sr, Ba, Ra

6. Order the following elements by increasing electronegativity:

B, Al, Ga, In, Tl

7. Order the following elements by increasing ionization energy:

Be, Mg, Ca, Sr, Ba, Ra

8. Order the following elements by increasing atomic radius:

Na, K, Rb, Cs, Fr

9. Order the following elements by increasing electronegativity:

Li, Na, K, Rb, Cs, Fr

10. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

Name: _____ Date: _____ Period: _____

Periodic Trends Test Study Guide

Section 1: Atomic Structure and Properties of the Periodic Table

1. What are the three main categories of elements in the periodic table?

- Metals
- Non-metals
- Metalloids

2. List the following elements in order of increasing atomic radius: Li, Na, K, Rb, Cs, Fr.

3. List the following elements in order of increasing electronegativity: F, Cl, Br, I, At.

4. List the following elements in order of increasing ionization energy: Li, Na, K, Rb, Cs, Fr.

5. List the following elements in order of increasing atomic radius: Be, Mg, Ca, Sr, Ba, Ra.

6. List the following elements in order of increasing electronegativity: B, Al, Ga, In, Tl.

7. List the following elements in order of increasing ionization energy: Be, Mg, Ca, Sr, Ba, Ra.

8. List the following elements in order of increasing atomic radius: Na, K, Rb, Cs, Fr.

9. List the following elements in order of increasing electronegativity: Li, Na, K, Rb, Cs, Fr.

10. List the following elements in order of increasing ionization energy: Li, Na, K, Rb, Cs, Fr.

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