

KEY

Percent, Actual, and Theoretical Yield



a) I began this reaction with 20 grams of lithium hydroxide. What is my theoretical yield of lithium chloride?

$$\frac{20 \text{ g LiOH}}{23.95 \text{ g LiOH}} \times \frac{1 \text{ mol LiOH}}{1 \text{ mol LiOH}} \times \frac{42.39 \text{ g LiCl}}{1 \text{ mol LiCl}} = \boxed{35.40 \text{ g LiCl}}$$

b) I actually produced 6 grams of lithium chloride. What is my percent yield?

$$\frac{6 \text{ g LiCl}}{35.40 \text{ g LiCl}} \times 100 = \boxed{16.9\%}$$



a) If I start with 5 grams of C_3H_8 , what is my theoretical yield of water?

$$\frac{5 \text{ g C}_3\text{H}_8}{44.1 \text{ g C}_3\text{H}_8} \times \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{8.175 \text{ g H}_2\text{O}}$$

b) I got a percent yield of 75%. How many grams of water did I get?

$$(8.175 \text{ g H}_2\text{O})(.75) = \boxed{6.13 \text{ g H}_2\text{O}}$$



My theoretical yield of barium chloride was 10.7 grams. If my actual yield was 9.5 grams, what was my percent yield?

$$\frac{9.5}{10.7} \times 100 = \boxed{88\%}$$



What is my theoretical yield of sodium oxide if I start with 20 grams of calcium oxide?

$$\frac{20 \text{ g CaO}}{56.08 \text{ g CaO}} \times \frac{1 \text{ mol CaO}}{1 \text{ mol CaO}} \times \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{22.10 \text{ g Na}_2\text{O}}$$

Percent Actual And Theoretical Yield Key

Rosina Ehmann



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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation *Introduction to General, Organic, and Biochemistry* Morris

Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career **EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS**

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